

# SUPPLEMENTARY INFORMATION

## Fluorinated materials as positive electrodes for Li- and Na-ion batteries

Kévin Lemoine<sup>a,\*</sup>, Annie Hémon-Ribaud<sup>a</sup>, Marc Leblanc<sup>a</sup>, Jérôme Lhoste<sup>a</sup>,  
Jean-Marie Tarascon<sup>b</sup> and Vincent Maisonneuve<sup>a,\*</sup>

<sup>a</sup>*Institut des Molécules et des Matériaux du Mans (IMMM) - UMR CNRS 6283, Le Mans Université, Avenue Olivier Messiaen, 72085 Le Mans Cedex 9, France*

<sup>b</sup>*Collège de France, Chaire de Chimie du Solide et de l'Energie, UMR 8260 CNRS, 11 Place Marcellin Berthelot, 75231 Paris, France ; Réseau sur le Stockage Electrochimique de l'Energie (RS2E), FR CNRS 3459, 80039 Amiens, France*

\*Corresponding authors : Vincent Maisonneuve, [vincent.maisonneuve@univ-lemans.fr](mailto:vincent.maisonneuve@univ-lemans.fr)  
Kevin Lemoine, [kevin.lemoine@uca.fr](mailto:kevin.lemoine@uca.fr)

This supplementary information brings together the main electrochemical performances of the fluorinated materials as a positive electrode in LIB and NIB configurations with the syntheses methods and the corresponding morphological characteristics. Fluorinated materials are classified along their structural dimensionality. Theoretical capacities are given along exchanged electron number related to insertion or conversion together with the potential window and the current density (nC means that the current will charge or discharge the full capacity in 1/n h).

Formulation	Synthesis method	Morphology, size	Theo. capacity (mAh.g <sup>-1</sup> )	Exp. capacity (mAh.g <sup>-1</sup> ) (cycle number)	Voltage range (V), Current density	Ref.	
<b>OD</b>							
<b>Cryolite structure type</b>							
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /C	Precipitation	Microparticles	140/1e <sup>-</sup>	116(1) / 90(20)	4.5-2.0, 0.2 mA.cm <sup>-2</sup>	1	
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /C	Precipitation	Nanoparticles (<50 nm)		100	4.5-2.0, 0.1 mA.cm <sup>-2</sup>	2,3	
$\beta$ -Li <sub>3</sub> FeF <sub>6</sub> /C	Precipitation	Nanoparticles (<50 nm)		80	4.5-2.0, 0.1 mA.cm <sup>-2</sup>	3	
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /C	Precipitation	Nanoparticles (<20 nm)		140(1) / 20(30)	4.0-2.0, C/18	4	
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /C	Sol-gel	Needles (13 nm)		110(1) / 70(100)	4.5-2.0, C/20, 7 mA.g <sup>-1</sup>	5	
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /CNT	Precipitation	Nanocomposites (100-300 nm)		120(1) / 75(50)	4.5-2.0, 14 mA.g <sup>-1</sup>	6	
$\alpha$ -Li <sub>3</sub> FeF <sub>6</sub> /Fe <sub>2</sub> O <sub>3</sub> /LiF/C coated	Precipitation	Nanocomposite (20 nm)		122(1) / 85(10)	4.0-2.0, C/50	7	
$\alpha$ -Li <sub>3</sub> VF <sub>6</sub> /AB	Solvothermal (microwaves)	Microparticles	144/1e <sup>-</sup>	110(1)	4.5-2.0, 0.2 mA.cm <sup>-2</sup>	8	
$\alpha$ -Li <sub>3</sub> Fe <sub>0.5</sub> V <sub>0.5</sub> F <sub>6</sub> /AB			142/1e <sup>-</sup>	90(1)			
$\beta$ -Li <sub>3</sub> VF <sub>6</sub> /AB			144/1e <sup>-</sup>	50(1)			
$\beta$ -Li <sub>3</sub> VF <sub>6</sub> /C	Precipitation	Nanoparticles (<20 nm)	144/1e <sup>-</sup>	144(1) / 100(10)	4.5-2.0, 0.1 mA.cm <sup>-2</sup>	9	
$\alpha$ -Li <sub>3</sub> CrF <sub>6</sub> /C	Sol-gel, heating	Nanoparticles (40 nm)	143/1e <sup>-</sup>	111(1) / 80(10)	4.5-1.5, C/20, 7 mA.g <sup>-1</sup>	10	
$\beta$ -Li <sub>3</sub> CrF <sub>6</sub> /C		Nanoparticles (200 nm)		106(1) / 90(10)			
Na <sub>3</sub> FeF <sub>6</sub> /Super P	NIB	Ball-milling	Nanoparticles (100 nm)	120(1) / 60(20)	4.25-0.5, C/10	11	
	LIB			500(1) / 200(20)			
Na <sub>3</sub> FeF <sub>6</sub> /C	Precipitation	Nanoparticles (200-500nm)	336/3e <sup>-</sup>	458(1) / 202(60)	4.5-1.0, 50 mA.g <sup>-1</sup>	12	
Na <sub>3</sub> FeF <sub>6</sub> /CNT	Precipitation	Nanoparticles (500 nm)		428(1) / 297(60)	4.5-1.0, 50 mA.g <sup>-1</sup>	12	
Na <sub>3</sub> FeF <sub>6</sub> /AB	NIB	Precipitation		Microparticles (1-3 $\mu$ m)	121(1) / 70(100)	4.0-0.4, 0.1C	13
	LIB				473(1) / 203(100)		
Na <sub>3</sub> FeF <sub>6</sub> /AB	NIB	Precipitation		Nanoparticles (150-250nm)	139(1) / 88(400)	4.0-0.5, 0.05C	14
	LIB				511(1) / 245(400)		
Na <sub>3</sub> FeF <sub>6</sub> /C	NIB	Microwave		Nanocomposite	185(1) / 45(20)	4.0-0.5, 0.05C	15
		Ball-milling	165(1) / 28(20)				
K <sub>3</sub> FeF <sub>6</sub> /C	Hydrothermal	Nanoparticles (30-50 nm)	279/3e <sup>-</sup>	213(1) / 140(30)	4.5-1.0, 0.2 mV.s <sup>-1</sup>	16	
Na <sub>3</sub> VOF <sub>5</sub> /C	Solvothermal	Microparticles	151/1e <sup>-</sup>	-	4.5-0.5, C/25	17	
<b>1D</b>							
<b>FeF<sub>3</sub>(H<sub>2</sub>O)<sub>2</sub>·H<sub>2</sub>O</b>							
FeF <sub>3</sub> (H <sub>2</sub> O) <sub>2</sub> ·H <sub>2</sub> O/AB	Precipitation	Nanocomposite with crystallites (100-1000 nm)	160/1e <sup>-</sup>	100(1) / 112(30)	4.5-2.0, 23.7 mA.g <sup>-1</sup>	18	

FeF <sub>3</sub> (H <sub>2</sub> O) <sub>2</sub> ·H <sub>2</sub> O/AB		Precipitation	Naocomposites with cristallites (6-13 nm)	481/3e <sup>-</sup>	342(1) / 111(35)	4.5-1.0, 0.1C	19
					92(1) / 83(100)	4.5-2.0, 0.1C	
FeF <sub>3</sub> (H <sub>2</sub> O) <sub>2</sub> ·H <sub>2</sub> O/CB		Precipitation	Aggregates (170 nm)		350(1) / 280(4)	4.5-1.0, 10 mA.g <sup>-1</sup>	20
				160(1) / 100(50)	4.5-2.0, 10 mA.g <sup>-1</sup>		
FeF <sub>3</sub> (H <sub>2</sub> O) <sub>2</sub> ·H <sub>2</sub> O/AB		Ball-milling	Nanoparticles (40-60 nm)		200(1) / 147(10)	4.5-1.8, 20 mA.g <sup>-1</sup>	21
<b>Na<sub>2</sub>MnF<sub>5</sub></b>							
Na <sub>2</sub> MnF <sub>5</sub> /C	NIB	Precipitation	Microparticles (10 μm)	137/1e <sup>-</sup>	103(1) / 50(2)	4.5-0.5, C/25	22
<b>2D</b>							
Ag <sub>2</sub> V <sub>4</sub> O <sub>11</sub> (SVO)		Solid	Microparticles	315/7e <sup>-</sup>	270(1) / -	3.5-1.0	23
Ag <sub>4</sub> V <sub>2</sub> O <sub>6</sub> F <sub>2</sub> (SVOF)		Hydrothermal	Microparticles	250/6e <sup>-</sup>	250(1) / -	3.5-0.5	24
Ag <sub>4</sub> V <sub>2</sub> O <sub>6</sub> F <sub>2</sub> (SVOF)		Precipitation	Microparticles (0.5 μm)		168(1) / -	3.5-0.5	25
AgNa(VO <sub>2</sub> F <sub>2</sub> ) <sub>2</sub>	NIB	Hydrothermal	Microparticles	432/6e <sup>-</sup>	125(1) / -	3.5-0.5	26
	LIB				90(1) / -		
<b>2D</b>							
<b>NaMF<sub>4</sub></b>							
NaVF <sub>4</sub> /C	NIB	Solvothermal	Microparticles	179/1e <sup>-</sup>	-	4.5-0.5, C/25	17
NaMnF <sub>4</sub> /C	NIB	Dehydration	Microparticles	174/1e <sup>-</sup>	-	4.5-0.5, C/25	22
NaFeF <sub>4</sub> /C	NIB	Solvothermal	Microparticles	173/1e <sup>-</sup>	12(1) / 10(3)	4.5-2.6, C/10	27
<b>Chiolite</b>							
Na <sub>5</sub> V <sub>3</sub> F <sub>14</sub>	NIB	Solvothermal	Microparticles	255/5e <sup>-</sup>	-	4.5-0.5, C/25	17
Na <sub>5</sub> Ti <sub>3</sub> O <sub>3</sub> F <sub>11</sub>	NIB	Solvothermal	Microparticles	260/5e <sup>-</sup>	-	4.5-0.5, C/25	
<b>3D</b>							
<b>Rutile</b>							
FeF <sub>2</sub>		Pulsed Laser Deposition	Thin films (thickness 140nm)	572/2e <sup>-</sup>	750(2) / 745(17)	3.6-0.05, 2.8 μA cm <sup>-2</sup>	28
FeF <sub>2</sub> /C		Precipitation, decomposition	Nanoparticles (10 nm)		480(1) / 480(5)	4.5-1.5, 4 mA g <sup>-1</sup>	29
FeF <sub>2</sub> /C		Reverse micro-emulsion	Micro and Nanorods		590(1) / 220(25)	5.0-1.0, 20 mA g <sup>-1</sup>	30
FeF <sub>2</sub> /C	NIB	Precipitation	Nanoparticles (10-20 nm)		190(1) / 70(6)	3.5-1.0, 10 mA g <sup>-1</sup>	31
FeF <sub>2</sub> @CNT (core/shell)		Solid	Nanorods (200-500nm x 2-4μm)		263(1) / 263(50)	4.2-1.0, 50 mA g <sup>-1</sup>	32
FeF <sub>2</sub> @C (core/shell)		Solid	Short rods (100nm-1μm x 1 μm)		314(1) / 217(50)	4.3-1.2, 30 mA g <sup>-1</sup>	33
Nanoconfined FeF <sub>2</sub> /C nanopores		Decomposition	Nanocomposite (100 nm)		120(1) / 100(200)	4.0-1.5, 150 mA g <sup>-1</sup>	34
FeF <sub>2</sub> @CMK-3		Fluorination	Nanoparticles		529(1) / 482(1000)	4.4-1.4, 500 mA g <sup>-1</sup>	35
FeF <sub>2</sub> /C		Precipitation/Decomposition	Nanocomposite (20 nm)		519(1) / 80(20)	4.5-1.5, 50 mA g <sup>-1</sup>	36

FeF <sub>2</sub> /C nanopores		Decomposition	Nanocomposite particles (1 μm)	572/2e <sup>-</sup>	450(1) / 300(1000)	4.0-1.5, 140 mA g <sup>-1</sup>	37	
FeF <sub>2</sub> /C	LIB	Precipitation/Decomposition	Nanoparticles (20-300 nm)		470(1) / 110(50)	4.2-1.5, 100 mA g <sup>-1</sup>	38	
	NIB				115(1) / 50(50)	4.0-1.2, 100 mA g <sup>-1</sup>		
Ni@FeF <sub>2</sub> @Al <sub>2</sub> O <sub>3</sub>		Electrodeposition/fluorination/atomic layer deposition	Open porous structure		473(1) / 200(100)	4.2-1.2, 200 mA g <sup>-1</sup>		39
FeF <sub>2</sub> /rGO-PAA (binder)	NIB	Precipitation/heating	Nanocomposite (10-20 nm)		135(1) / 120(120)	4.5-1.5, 200 mA g <sup>-1</sup>	40	
FeF <sub>2</sub> @CNT/SPE		Solid	Nanoparticles (10-20 nm)		700(1) / 450(100)	4.0-1.0, 50 mA g <sup>-1</sup>	41	
FeF <sub>2</sub> /C		Catalytic decomposition PFPE	Nanoparticles (10 nm)		500(1) – 50(20)	4.5-1.0 – C/20	42	
FeF <sub>2</sub> @MHCS/C	NIB	Infiltration/heating	Nanocomposite particles (4-20 nm)		134(1) / 112(100)	4.5-1.5, 200 mA g <sup>-1</sup>	43	
FeF <sub>2</sub> @GC (MOF)	NIB	Carbonization/Precipitation/Solvothermal	Nanocomposite (10-15 nm)		611(1) / 120(300)	4.5-0.8, 50 mA g <sup>-1</sup>	44	
FeF <sub>2</sub> /C/LP30(electrolyte) (RT)		Colloidal synthesis	Nanorods (20 nm)		755(1) / 0(50)	4.0-1.0, C/20	45	
FeF <sub>2</sub> /C/IL(electrolyte) (50°C)					700(1) / 525(50)	4.0-1.0, C/20		
FeF <sub>2</sub> @rGO/C		fluorination	Porous composite		430(1) / 400(50)	4.0-1.0, 80 mA g <sup>-1</sup>	46	
NiF <sub>2</sub>		Pulsed Laser Deposition	Thin Film		554/2e <sup>-</sup>	650(1) / 465(100)	3.5-0.01, 10 μA.cm <sup>-2</sup>	47
NiF <sub>2</sub> /C		Ball-milling	Nanoparticles			1100(1) / 200(10)	4.0-0.5, 5.5 mA g <sup>-1</sup>	48
NiF <sub>2</sub> /C		Heating	-	740(1) / 700(2)		4.5-1.0, 16 mA g <sup>-1</sup>	49	
MnF <sub>2</sub> /C		Solvothermal/IL	Nanoparticles (100-300 nm)	576/2e <sup>-</sup>	725(1) / 237(5000)	3.0-0.01, 10C	50	
MnF <sub>2</sub>		Solvothermal/IL	3D nanorods (20x200 nm)		960(1) / 420(2000)	3.0-0.01, 10C	51	
MnF <sub>2</sub> /graphene		Solvothermal/IL	Nanoparticles (50-200 nm)		250(1) / 290(4000)	3.0-0.01, 6 A g <sup>-1</sup>	52	
MnF <sub>2</sub> /C	LIB	Precipitation/Decomposition	Nanoparticles (20-300 nm)		100(1) / 55(100)	4.2-1.5, 100 mA g <sup>-1</sup>	38	
	NIB				100(1) / 50(100)	4.0-1.2, 100 mA g <sup>-1</sup>		
MnF <sub>2</sub> /MWCNT		Precipitation/decomposition	Hierarchical composite (140 nm)		665(1) / 480(100)	4.0-0.5, 5.5 mA g <sup>-1</sup>	53	
CoF <sub>2</sub>		Pulsed Laser Deposition	Thin Film	552/2e <sup>-</sup>	600(1) / 160(8)	3.5-0.01, 10 μA.cm <sup>-2</sup>	54	
CoF <sub>2</sub> /MWCNT(NC3100)		Precipitation/heating	Nanoparticles (20 nm)		560(1) / 320(10)	4.3-1.0, 50 mA g <sup>-1</sup>	55	
CoF <sub>2</sub> /C		Reverse micro-emulsion	Microspheres (0.7-1.8 μm)		535(1) / 30(25)	4.8-1.2, 20 mA g <sup>-1</sup>	30	
CoF <sub>2</sub> /AB	LIB	Precipitation/heating	Nanoparticles (10 nm)		858(1) / 166(100)	3.2-0.01, 111 mA g <sup>-1</sup>	56	
	NIB			402(1) / 47(30)	3.0-0.01, 553 mA g <sup>-1</sup>			

CoF <sub>2</sub> @CNT		CVD/Infiltration/Decomposition	Nanocomposite (30-50 nm)	552/2e <sup>-</sup>	360(1) / 335(200)	4.0-1.0, 100 mA g <sup>-1</sup>	57
CoF <sub>2</sub> /C		Solvothermal	Porous nanospheres (150-500 nm)		538(1) / 127(30)	4.8-1.0, 20 mA g <sup>-1</sup>	58
CoF <sub>2</sub> /C	LIB	Precipitation/Decomposition	Nanoparticles (20-300 nm)		280(1) / 70(50)	4.2-1.5, 100 mA g <sup>-1</sup>	38
	NIB				100(1) / 65(50)	4.0-1.2, 100 mA g <sup>-1</sup>	
CoF <sub>2</sub> @CFC		Solvothermal/IL	Composite (30-40 nm)		685(1) / 330(200)	4.0-1.0, 100 mA g <sup>-1</sup>	59
CoF <sub>2</sub> /Fe <sub>2</sub> O <sub>3</sub> /C (Fe-Co-ZIF)		Precipitation/Pyrolysis	Nanoparticles (500 nm)		387(1) / 198(100)	4.5-1.2, 50 mA g <sup>-1</sup>	60
CoF <sub>2</sub> @C (MOF-67)		Carbonization/Fluorination	MOF-67 shaped nanocomposite (5-20 nm)		472(1) / 380(200)	4.0-1.0, 0.2C	61
CuF <sub>2</sub> /C		Ball-milling	Nanoparticles	528/2e <sup>-</sup>	303(1)	4.0-2.0, C/35	62
CuF <sub>2</sub> /C		Solvothermal	Nano (50-100 nm)		350(1) / 40(3)	4.5-2.0, 0.1C	63
Ni <sub>0.75</sub> Co <sub>0.25</sub> F <sub>2</sub> /		Precipitation/heating	Nanoparticles (16-22 nm)	553/2e <sup>-</sup>	700(1) / 460(10)	4.3-1.0, 50 mA g <sup>-1</sup>	64
Fe <sub>0.9</sub> Ni <sub>0.1</sub> F <sub>2</sub> /MWCNT		Precipitation/decomposition	Nanoparticles (100 nm)	570/2e <sup>-</sup>	425(1) / 175(50)	4.0-1.0, 140 mA g <sup>-1</sup>	65
Fe <sub>0.5</sub> Cu <sub>0.5</sub> F <sub>2</sub> /C		Ball-milling	Nanoparticles	547/2e <sup>-</sup>	575(1) / 480(5)	4.5-1.0, 9.2 mA g <sup>-1</sup>	66
NiF <sub>2</sub> /C		Precipitation/Decomposition	Nanoparticles (20 nm)	554/2e <sup>-</sup>	445(1) / 40(6)	4.0-1.0, C/10	67
Cu <sub>0.1</sub> Ni <sub>0.9</sub> F <sub>2</sub> /C			Nanoparticles (20 nm)	553/2e <sup>-</sup>	548(1) / 155(6)		
Cu <sub>0.25</sub> Ni <sub>0.75</sub> F <sub>2</sub> /C			Nanoparticles (50-150 nm)	547/2e <sup>-</sup>	552(1) / 220(6)		
CuF <sub>2</sub> /MoO <sub>3</sub> /C		Ball-milling	Nanoparticles (2-30 nm)	528/2e <sup>-</sup>	450(1)	3.5-2.0, 7.58 mA g <sup>-1</sup>	68
FeOF/C		Precipitation/heating	Nanoparticles (7-15 nm)	885/3e <sup>-</sup>	410(1) / 280(60)	4.5-1.5, 50 mA g <sup>-1</sup> (60°C)	69
FeOF/C		<i>Roll quenching</i>	-		900(1) / 40(10)	4.0-0.7, 10 mA g <sup>-1</sup>	70
FeOF/C					560(1) / 360(30)	4.0-1.3, 10 mA g <sup>-1</sup>	
FeOF/C		Solvothermal	Nanorods (<1 μm)		410(1) / 340(20)	4.5-1.5, 50 mA g <sup>-1</sup> (60°C)	71
FeOF@PEDOT		Solvothermal/Polymerization	Core-shell Nanorods (<1 μm)		560(1) / 430(150)	4.2-1.2, 50 mA g <sup>-1</sup>	72
FeOF@PEDOT		Solvothermal/Polymerization	Coated Nanorods (30-50nm x 1 μm)		165(1) / 125(200)	3.8-2.0, 100 mA g <sup>-1</sup> (60°C)	73
FeOF/FeF <sub>3</sub> /C		Ball-milling/oxidation process	Nanoparticles		438(1) / 130(50)	4.5-1.5, 20 mA g <sup>-1</sup>	74
FeOF/rGO		Solvothermal/reduction process	Nanorods (30 x 500 nm)	400(1) / 327(100)	4.5-1.5, 100 mA g <sup>-1</sup>	75	
LiF/FeO composite/C (Cubic FeOF)		Ball-milling	Nanoparticles (20-50 nm)	295/1e <sup>-</sup>	225(1) / 275(30)	4.8-1.5, 50 mA g <sup>-1</sup> (60°C)	76
FeO <sub>0.7</sub> F <sub>1.3</sub> /C	NIB	Precipitation	Nanocomposite (15 nm)	885/3e <sup>-</sup>	496(1) / 360(50)	3.5 -1.0, 25 mA g <sup>-1</sup> (50°C)	77
FeOF/C	NIB	Solvothermal	Hierarchical Nanorods (20-35 nm)		450(1) / 130(70)	4.0-1.0, 20 mA g <sup>-1</sup>	78

FeOF/C	NIB	Solution Plasma Processing	Mesoporous amorphous Nanococoons (75x25 nm)	885/3e <sup>-</sup>	290(1) / 270(100)	3.8-1.3, 1 mA g <sup>-1</sup> (50°C)	79
FeOF/C	NIB	<i>Roll quenching</i>	-		240(1) / 210(20)	4.0-1.0, 10 mA g <sup>-1</sup>	80
FeOF/rGO	NIB	Solvothermal/Thermal reduction	Nanoparticles (20-35 x 80-110 nm)		283(1) / 220(100)	4.0-1.2, 20 mA g <sup>-1</sup>	81
FeOF/GCL (MOF)	NIB	Solvothermal/Oxydation process	Nanoparticles (50 nm)		437(1) / 338(100)	4.0-1.2, 100 mA g <sup>-1</sup>	82
FeOF/GC (rosin acid)	NIB	Solvothermal/Oxydation process	Wrapped Nanospheres (10-20 nm)		449(1) / 357	4.0-1.2, 100 mA g <sup>-1</sup>	83
Co <sub>0.1</sub> Fe <sub>0.9</sub> OF		Solvothermal	Nano (50-100 nm)		550(1) / 350(1000)	4.0-1.2, 500 mA g <sup>-1</sup>	84
<b>Dirutile</b>							
LiMnF <sub>4</sub>		Ball-milling	Nano (100-200 nm)	194/1e <sup>-</sup>	30(1) / 45(2)	4.5-2.0, C/50	85
<b>Trirutile</b>							
LiFe <sub>2</sub> F <sub>6</sub> /C		Ball-milling	Nanoparticles	230/2e <sup>-</sup>	126(1) / 78(14)	4.5-2.0, C/12	86
Li <sub>1.2</sub> Fe <sub>2</sub> F <sub>6</sub> /C		Ball-milling	Nanoparticles	228/2e <sup>-</sup>	155(1) / 137(27)	4.5-2.0, C/12	86
LiMgFeF <sub>6</sub> /C		Sol-gel/heating	Nano (100-200 nm)	133/1e <sup>-</sup>	90(1) / 110(20)	4.5-2.0, C/20	87
LiNiFeF <sub>6</sub> /C		Sol-gel/heating	Nano (100-200 nm)	114/1e <sup>-</sup>	95(1) / 88(20)	4.5-2.0, C/20	88
Li <sub>2</sub> TiF <sub>6</sub> /C		Atomization	Nano (500 nm)	152/1e <sup>-</sup>	110(1) / 80(16)	4.5-2.0, 0.2 mA g <sup>-1</sup>	89
<b>Colquiriite</b>							
LiCaFeF <sub>6</sub> /C		Solid/Ball-milling	Nanoparticles (240 nm)	124/1e <sup>-</sup>	112(1) / 95(20)	4.5-2.0, C/20	90
<b>Na<sub>2</sub>SiF<sub>6</sub></b>							
LiMnFeF <sub>6</sub> /C		Sol-gel/heating	Nano (250 nm)	116/1e <sup>-</sup>	95(1) / 73(10)	4.3-2.2, C/20	91a
HP-LiFe <sub>2</sub> F <sub>6</sub> /AB		High-pressure	Microparticles	116/1e <sup>-</sup>	155(1) / 102(10)	4.5-2.0, C/20	91b
<b>Inverse Spinel</b>							
Li <sub>2</sub> NiF <sub>4</sub> /C		Sol-gel	Microparticles	180/1e <sup>-</sup>	750(1) / 31(20)	3.8-0.6, 9 mA g <sup>-1</sup>	92
Li <sub>2</sub> NiF <sub>4</sub> /PEDOT		Solvothermal	Nanocomposite		550(1) / 303(40)	3.8-0.5, 10 mA g <sup>-1</sup>	93
Li <sub>2</sub> NiF <sub>4</sub> /C		ball-milling / calcination	Microparticles?		96(1)	4.8-2.0, C/20	94
2LiF-NiO/C		Ball-milling	Nanocomposite	-	146(1)	4.8-2.0, C/10	94
MnO-LiPF <sub>6</sub>		<i>In situ</i>	-	-	330(1) / 220(5)	4.5-1.5, C/50	95,96
<b>BiF<sub>3</sub>/BiOF</b>							
BiF <sub>3</sub> /C		Ball-milling	Nanoparticles	302/3e <sup>-</sup>	225(1) / 200(15)	4.5-2.0, 45 mA g <sup>-1</sup>	97,98
BiF <sub>3</sub> /C (55°C)		Ball-milling	-		314(1) / 183(10)	4.5-2.0, 0.1C	99
BiOF/C		Ball-milling	Nanoparticles	330/3e <sup>-</sup>	180(1) / 100(5)	4.5-1.8, 23 mA g <sup>-1</sup>	100
BiO <sub>0.5</sub> F <sub>2</sub> /C		Ball-milling	Nanoparticles	315/3e <sup>-</sup>	250(1) / 187(3)	4.5-2.0, 7.6 mA g <sup>-1</sup>	100
BiO <sub>0.5</sub> F <sub>2</sub> @CMK-3		Infiltration/Heating	Nanoparticles (20 nm) composite		343(1) / 148(40)	4.5-1.5, 30 mA g <sup>-1</sup>	101

Weberite								
Na <sub>2</sub> Fe <sub>2</sub> F <sub>7</sub> /C	NIB	Solvothermal	Nanoparticles (500 nm)	184/2e <sup>-</sup>	58(1) / 50(30)	3.8-2.6, 0.1C	27	
Na <sub>2</sub> Fe <sub>2</sub> F <sub>7</sub> /C	NIB	Ball-milling/heating	Nanoparticles (25-250 nm)		135(1) / 115(1000)	4.3-1.5, 2C	102	
Na <sub>2</sub> FeVF <sub>7</sub> /C	NIB	Ball-milling/heating	Nanoparticles (200 nm)	188/2e <sup>-</sup>	105(1) / 100(200)	4.5-1.5, 100 mA g <sup>-1</sup>	103	
Na <sub>2</sub> MnVF <sub>7</sub> /C	NIB	Ball-milling/heating	Nanoparticles (200 nm)	146/2e <sup>-</sup>	85(1) / 80(200)	4.5-1.5, 100 mA g <sup>-1</sup>	103	
Na <sub>2</sub> CoVF <sub>7</sub> /C	NIB	Ball-milling/heating	Nanoparticles (200 nm)	145/2e <sup>-</sup>	65(1) / 60(200)	4.5-1.5, 100 mA g <sup>-1</sup>	103	
Perovskite								
NaFeF <sub>3</sub> /C	NIB	Ball-milling	Agglomerats (6 μm)	197/1e <sup>-</sup>	125(1) / 100(20)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	104	
NaFeF <sub>3</sub> /C	NIB	Precipitation	Nanoparticles (500 nm)		253(1)	4.0-1.5, 0.1C	105	
NaFeF <sub>3</sub> /C	NIB	Roll-quenching	Nanoparticles (10-20 nm)		200(1)	4.5-1.5, 0.014C	106	
NaFeF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (30 nm)		200(1) / 250(10)	4.5-1.5, 7.5 mA g <sup>-1</sup>	107	
NaFeF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (20-100nm)		169(1) / 100(20)	4.3-2.0, 0.1 mA cm <sup>-2</sup>	108	
NaFeF <sub>3</sub> /C	NIB	Solvo. microwaves	Nanoparticles (500 nm)		200(1) / 140(60)	4.0-2.0, 0.1C	109	
	LIB				200(1) / 200(60)	4.0-2.0, 0.33C		
NaFeF <sub>3</sub> /C	NIB	Solvothermal	Nanoparticles (200 nm)		153(1) / 75(200)	4.5-1.5, 1C	110	
	LIB				183(1) / 100(200)			
NaFeF <sub>3</sub> /C	NIB	Precipitation	Nanoparticles (100 nm)		154(1) / 119(400)	4.0-1.5, 0.05C	14	
	LIB				181(1) / 158(400)	4.2-2.0, 0.05C		
NaCuF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (30 nm)		372/2e <sup>-</sup>	200(1) / 50(7)	4.5-1.5, 7.5 mA g <sup>-1</sup>	107
NaNiF <sub>3</sub> /C	NIB	Ball-milling	Agglomerats (7.6 μm)		193/1e <sup>-</sup>	30(1) / 42(2)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	104
NaNiF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (30 nm)			40(1) / 35(1)	4.5-1.5, 7.5 mA g <sup>-1</sup>	107
NaCoF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (30 nm)	193/1e <sup>-</sup>	45(1) / 40(2)	4.5-1.5, 7.5 mA g <sup>-1</sup>	107	
NaCoF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (20-100nm)		38 (1) / -	4.6-2.0, 0.1 mA cm <sup>-2</sup>	108	
NaCoF <sub>3</sub> /rGO		Solvothermal	Nanoparticles (20 nm)/nanocluster (300-500 nm)		514 (1) / 350 (5)	4.0-1.0, 20 mA g <sup>-1</sup>	111	
NaMnF <sub>3</sub> /C	NIB	Ball-milling	Agglomerats (10 μm)	198/1e <sup>-</sup>	32(1) / 37(2)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	104	
NaMnF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (30nm)		50(1) / 48(2)	4.5-1.5, 7.5 mA g <sup>-1</sup>	107	
NaMnF <sub>3</sub> /AB	NIB	Ball-milling	Nanoparticles (20-100nm)		89(1) / 40(20)	4.6-2.0, 0.1 mA cm <sup>-2</sup>	108	
KFeF <sub>3</sub> /C	NIB	Precipitation	Nanoparticles (100 nm)	180/1e <sup>-</sup>	170(1) / 110(35)	4.5-1.5, 0.1C	112	
KFeF <sub>3</sub> /C	LIB	Ball-milling	Microparticles (1μm)		140(1) / 80(80)	4.5-1.8, C/30	113	
	NIB				120(1) / 90(30)	4.2-1.8, C/30		
AgCuF <sub>3</sub> /C		Ball-milling	Nanoparticles (20-40nm)	234/2e <sup>-</sup>	264(1)	4.0-2.0, 7.58 mA g <sup>-1</sup>	114	
FeF <sub>3</sub> /C (50:50 wt %)		Ball-milling	Nanodomains (25-30 nm)	237/1e <sup>-</sup>	200(1) / 140(10)	4.5-2.5, 7.58 mA g <sup>-1</sup>	115	



FeF <sub>3</sub> /CB (85:15 wt %)		Ball-milling	Nanoparticles (34 nm)	712/3e <sup>-</sup>	660 (2) / 600(12)	4.5-1.5, 7.58 mA g <sup>-1</sup> (70°C)	116
FeF <sub>3</sub> /C (75:25 wt %)		Ball-milling ( <i>In situ</i> redox)	Nanoparticles (20 nm)		380(1) / 315(12)	4.5-1.5, 7.58 mA g <sup>-1</sup>	
FeF <sub>3</sub> /AB	LIB	Ball-milling	Nanoparticles	237/1e <sup>-</sup>	210(1) / 150(40)	4.0-2.0, 0.2 mA cm <sup>-2</sup>	118
	NIB				150(1) / 90(10)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	
(FeF <sub>3</sub> /CNT)/C		Precipitation, heating	FeF <sub>3</sub> nanoflowers (10nm) on CNTs	237/1e <sup>-</sup>	210(1) / 200(30)	4.5-2.0, 20 mA.g <sup>-1</sup>	119
FeF <sub>3</sub> /AB	Precipitation with PEG, heating		Nanoparticles (15 nm), no agglomerates	712/3e <sup>-</sup>	860(1) / 742(10)	4.5-1.0, 100 mA.g <sup>-1</sup>	120
	Precipitation with CTAB, heating		Nanoparticles (10 nm)		615(1) / 615(10)		
	Solvothelmal, heating		Nanoparticles (10-20 nm)		550(1) / 547(10)		
FeF <sub>3</sub> /AB		Ball-milling, heating	Nanoparticles (30nm)	237/1e <sup>-</sup>	210(1) / 170(55)	4.5-2.0, 10 mA.g <sup>-1</sup>	121
FeF <sub>3</sub> /ΣFeF <sub>2</sub> /C		Precipitation, thermal dehydration	Porous nanowires (15µm×30-90nm)	712/3e <sup>-</sup>	543(1) / 100(100)	4.5-1.5, 50 mA.g <sup>-1</sup>	122
FeF <sub>3</sub> /CB		Ball milling	Nanoparticles (80-160nm)		710(1) / 284(100)	4.5-1.5, 21 mA.g <sup>-1</sup>	123
FeF <sub>3</sub> @Fe <sub>3</sub> O <sub>4</sub> core-shell		Sol-gel, fluorination, heating	Core (FeF <sub>3</sub> , 100-150nm)-shell (Fe <sub>3</sub> O <sub>4</sub> , 5nm) composite	237/1e <sup>-</sup>	200(1) / 80(12)	4.5-2.0, 50 mA.g <sup>-1</sup>	124
FeF <sub>3</sub> /AB		Solid state reaction(PTFE)	Nanoparticles (20-100 nm)	712/3e <sup>-</sup>	387(1) / 320(5)	4.5-1.0, 0.1C	125
FeF <sub>3</sub> /AB		Thermal evaporation and decomposition	Nanoparticles (30 nm)	237/1e <sup>-</sup>	224(1) / 143(100)	4.5-2.0, 0.1C	126
(FeF <sub>3</sub> /Graphene)/CB		Vapour solid autoclave method, heating	Rectangular rods (300×150 nm) wrapped by multi-sheet graphene		245(1) / 186(100)	4.5-1.5, 21 mA.g <sup>-1</sup>	127
(FeF <sub>3</sub> /C composite)/no CB		Vapour solid autoclave method, heating	FeF <sub>3</sub> rods (150-400nm) into porous carbon matrix	712/3e <sup>-</sup>	610(1) / 197(50)	4.5-1.5, 21 mA.g <sup>-1</sup>	128
FeF <sub>3</sub> /rGO		Precipitation, thermal dehydration	Nanocomposite, nanoparticles (70 nm)	237/1e <sup>-</sup>	150(1) / 135(50)	4.5-1.7, 500 mA.g <sup>-1</sup>	129
FeF <sub>3</sub> -graphene/C		Precipitation, thermal dehydration	Nanoparticles (20-100 nm) on graphene sheets		75(1) / 67(50)	4.5-2.0, 10C	130
FeF <sub>3</sub> /OMC		Solid template method (CCT)	Composite, nanoparticles (12, 23 nm)	237/1e <sup>-</sup>	163(1) / 135(30)	4.5-2.0, 0.1C	131
FeF <sub>3</sub> /C		Precipitation	Nanoparticles (5-10 nm)	712/3e <sup>-</sup>	606(1) / 50(35)	4.5-1.0, 0.1C	19
					170(1) / 110(100)	4.5-2.0, 0.1C	
(FeF <sub>3</sub> /Fe/rGO)/AB	NIB	Freeze drying, heating, electrochemical activation from FeF <sub>2</sub> -rGO	FeF <sub>2</sub> particles (20-40nm) on the rGO sheet	-	125(1) / 70(1000)	4.5-2.0, 100 mA.g <sup>-1</sup>	132
(FeF <sub>3</sub> /graphene)/C	NIB	Precipitation	Nano-sized composite (2-30nm)	237/1e <sup>-</sup>	210(1) / 106(50)	4.2-1.5, 60 mA.g <sup>-1</sup>	133
FeF <sub>3</sub> /printex C		Ball-milling assisted low-temperature	Nanoparticles (10-20nm) in agglomerats (200×100nm)		225(1) / 150(60)	4.5-2.0, 100 mA.g <sup>-1</sup>	134
(FeF <sub>3</sub> /graphene)/ C					235(1) / 155(60)		

(FeF <sub>3</sub> /amorphous C)/C		low-temperature precipitation, heating	Nanocomposite, FeF <sub>3</sub> particles (20-30nm)		224(1) / 130(60)			
FeF <sub>3</sub> /CMB		Fluorination	Porous hybrid composite, nanocrystals (1-4 nm)	712/3e <sup>-</sup>	450(1) / 336(500)	4.5-1.5, 1 A.g <sup>-1</sup>	135	
(FeF <sub>3</sub> /graphene)/SP	NIB	Sol-gel method, heating	Composite (FeF <sub>3</sub> nanosheets +graphene sheets (100 nm->10µm))		344(1) / 116(50)	4.0-1.0, 0.3C	136	
(FeF <sub>3</sub> /rGO)/CB		Hydrothermal, heating dehydration	Composite, FeF <sub>3</sub> submicron microspheres with particles (30-40nm)	237/1e <sup>-</sup>	196(1) / 169(50)	4.5-2.0, 0.1C	137	
(FeF <sub>3</sub> /mesoporous ACF)/CB		Impregnation, heating	FeF <sub>3</sub> nanoparticles (5-6nm) inside pores (6nm) of mesoporous carbon		195(1) / 181(50)	4.5-2.0, 0.1C	138	
FeF <sub>3</sub> @NAN/AC		Encapsulation, heating	Core (FeF <sub>3</sub> nanospheres (8 nm)) -shell (NAN) hybrid architecture	712/3e <sup>-</sup>	426(1) / 328(400)	4.5-1.0, 71 mA.g <sup>-1</sup>	139	
FeF <sub>3</sub> /graphitic carbon/CB		Polymerization, heating, fluorination, dehydration heating	Nanocomposite (worm-like nanoparticles (10-20 nm) wrapped by carbon matrix (5-7 nm))	237/1e <sup>-</sup>	188(1) / 166(50)	4.5-2.0, 23.7 mA.g <sup>-1</sup>	140	
Fe/LiF/CB		Ball-milling	Nanocomposites (Fe, Li nanoparticles (6 nm) and C)	712/3e <sup>-</sup>	430(1) / 310(50)	4.5-1.2, 225 mA.g <sup>-1</sup>	141	
FeF <sub>3</sub> /AB		F <sub>2</sub> fluorination	Low crystallinity (pore size 3 nm)		676(1) / 95(15)	4.5-1.0, 71 mA.g <sup>-1</sup>	142	
FeF <sub>3</sub> /amorphous C		Solid state reaction, decomposition	Nanoparticles (10-40 nm) covered by amorphous carbon	237/1e <sup>-</sup>	250(1) / 127(100)	4.5-2.0, 50 mA.g <sup>-1</sup>	143	
FeF <sub>3</sub> /AC		Ball-milling, heating	Boron-based additives in electrolyte	712/3e <sup>-</sup>	624(2) / 540(10) -	4.5-1.0, 0.05C	144	
FeF <sub>3</sub> /AC		Gaseous fluorination 350°C	hierarchized macroporous/ mesoporous texture		317(1) / 80(20)	4.3-2.5, 0.05C, 60°C	145	
FeF <sub>3</sub> /rGO/C	LIB	Thermal decomposition	Nanoparticles (5-20 nm)	237/1e <sup>-</sup>	220(1) / 152(100)	4.2-2.0, 50 mA g <sup>-1</sup>	38	
	NIB				525(1) / 360(20)	4.2-1.5, 0.2 A g <sup>-1</sup>		
					160(1) / 145(100)	4.0-1.6, 0.2 A g <sup>-1</sup>		
FeF <sub>3</sub> /Super P		Thermal decomposition of (NH <sub>4</sub> ) <sub>3</sub> FeF <sub>6</sub> nanocrystals	Porous agglomerates of nanoparticles (5-20 nm)		220(1) / 188(50)	4.5-2.0, 0.2C	146	
FeF <sub>3</sub> /PHCNF interlayer		Solvothermal, thermal dehydration	Spindle-shaped agglomerates (10-50 µm) with nanoparticles (20 nm)			254(1) / 217(40)	4.5-2.0, 20 mA.g <sup>-1</sup>	147
FeF <sub>3</sub> /AB		Ball-milling	-		712/3e <sup>-</sup>	566(1) / 20(50)	4.5-1.0, 0.2 mA cm <sup>-2</sup>	148
					180(1) / 130(50)	4.5-2.0, 0.2 mA cm <sup>-2</sup>		
FeF <sub>3</sub> /AC		Solvothermal, thermal dehydration	Amorphous and crystallized fusiform porous particles (>10 µm x 20 nm)	237/1e <sup>-</sup>	125(1) / 137(100)	4.5-2.0, 20 mA.g <sup>-1</sup>	149	
FeF <sub>3</sub> /C		Solvothermal	Nanocomposite (carbon wrapped FeF <sub>3</sub> particles 50-60 nm)		166(1) / 126(100)	4.5-2.0, 20 mA.g <sup>-1</sup>	150	
FeF <sub>3</sub> /AC		Ball-milling	Nanocomposite (Particles (10-30 nm))	712/3e <sup>-</sup>	346(1) /161(40)	4.0-1.0, 50 mA.g <sup>-1</sup>	151	

FeF <sub>3</sub> /C nanofibers		Electrospinning, carbonization/reduction, fluorination (NF <sub>3</sub> )	Nanocomposite (nanofibers (>10 μm x 300-500 nm), nanoparticles (20-50 nm))		500(1) / 500(400)	4.0-1.0, 100 mA.g <sup>-1</sup>	152
FeF <sub>3</sub> /MWCNT//Li-B		Precipitation, thermal dehydration	Microcrystals (1-3x5-10 μm)	237/1e <sup>-</sup>	160 (500°C)	3.27 -100 mA.m <sup>-2</sup>	153
FeF <sub>3</sub> /C		Ball-milling	Nanoparticles	712/3e <sup>-</sup>	600(1) / 372(30)	4.5-1.0, 0.05C	154,155
FeF <sub>3</sub> /C		Ball-milling	Nanoparticles (13 nm)		610(1) / 380(30)	4.5-1.0, 0.1C	156
FeF <sub>3</sub> /3D honeycomb carbon framework		Sol-gel, carbonization, fluorination (Ar/NF <sub>3</sub> )	Nanospheres (10-50 nm), composite	237/1e <sup>-</sup>	211(1) / 64(1000)	4.5-1.0, 5C	157
FeF <sub>3</sub> /C		Catalytic decomposition PFPE	Nanoparticles (5-10 nm)		210(1) / 147(50)	4.5-1.0, C/20	42
FeF <sub>3</sub> /C		Solid gas F <sub>2</sub> reaction	Microcrystalline powder	712/3e <sup>-</sup>	200(1e <sup>-</sup> ), 700 (3e <sup>-</sup> )	3.5-1.0, 20 mA.g <sup>-1</sup>	158
FeF <sub>3</sub> /C nanocages		Solvothermal, thermal dehydration	Porous Nanocomposite (nanocrystals, 3-9 nm)		610(1) / 410(120)	4.5-1.7, 100 mA.g <sup>-1</sup>	159
Fe <sub>0.99</sub> Ti <sub>0.01</sub> F <sub>3</sub> /C		Hydrothermal/heating/ ball-milling	Nanocomposite (20-100 nm)	237/1e <sup>-</sup>	194(1) / 174(30)	4.5-2.5, 0.1C	160
FeF <sub>3</sub> /MWCNT		Solvothermal/heating	Nanocomposite (50-200 nm)		192(1) / 161(50)	4.5-2.0, 0.2C	161
Fe <sub>0.96</sub> Co <sub>0.04</sub> F <sub>3</sub> /MWCNT					217(1) / 188(50)		
FeF <sub>3</sub> /C		Pyrolysis	Porous spheres (100nm) with nanocages (5 nm)		145(1) / 60(100)	4.5-2.0, 50 mA.g <sup>-1</sup>	162
Fe <sub>0.99</sub> Co <sub>0.01</sub> F <sub>3</sub> /C					175(1) / 114(100)		
Fe <sub>0.97</sub> Co <sub>0.03</sub> F <sub>3</sub> /C				175(1) / 80(100)			
MnF <sub>3</sub> /AB	LIB	Ball-milling	Nanoparticles	239/1e <sup>-</sup>	52(1) / 25(2)	4.5-2.0, 0.2 mA cm <sup>-2</sup>	118
	NIB				72(1) / 18(2)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	
CoF <sub>3</sub> /C		Fluorination (100°C)	Nanoparticles (11 nm)	693/3e <sup>-</sup>	1011(1) / 400(14)	4.5-0.02, 5 mA g <sup>-1</sup>	163
CoF <sub>3</sub> /AB	LIB	Ball-milling	Nanoparticles	231/1e <sup>-</sup>	25 (1) / 10(2)	4.5-2.0, 0.2 mA cm <sup>-2</sup>	118
	NIB				38(1) / 10(2)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	
TiF <sub>3</sub> /C		Ball-milling	Nanoparticles (17 nm)	768/3e <sup>-</sup>	730(1) / 400(40)	4.0-0.5, 38 mA g <sup>-1</sup>	148,164
TiF <sub>3</sub> /C		-	Commercial product (μm)		940(1) / 540(15)	3.5-0.02, 0.2 mA cm <sup>-2</sup>	165
TiF <sub>3</sub> /AB	LIB	Ball-milling	Nanoparticles	256/1e <sup>-</sup>	62 (1) / 52(2)	4.5-2.0, 0.2 mA cm <sup>-2</sup>	118
	NIB				62(1) / 38(2)	4.0-1.5, 0.2 mA cm <sup>-2</sup>	
VF <sub>3</sub> /C		-	Commercial product (μm)	744/3e <sup>-</sup>	910(1) / 500(10)	4.3-0.02, 0.2 mA cm <sup>-2</sup>	165
AlF <sub>3</sub> /C		Ball-milling	Nanoparticles (50-600 nm)	957/3e <sup>-</sup>	790(1) / 300(12)	4.0-0.5, 0.1C	166
VO <sub>2</sub> F/C		High pressure solid state	Microparticles	788/3e <sup>-</sup>	410(1) / 200(14)	3.9-2.0, C/50	167,168
VO <sub>2</sub> F/C		Ball-milling	Nanoparticles		450(1) / 215(54)	4.1-1.3, 50 mA g <sup>-1</sup>	169
VO <sub>2</sub> F/graphene		Ball-milling	Nanoparticles		250(1) / 150(50)	4.3-2.1, C/20	170

Li <sub>2</sub> VO <sub>2</sub> F/C	Ball-milling	Nanoparticles	525/2e <sup>-</sup>	310(1) / 250(14)	4.5-1.3, C/20	171	
Li <sub>2</sub> CrO <sub>2</sub> F/C	Ball-milling	Nanoparticles	521/2e <sup>-</sup>	375(1) / 180(60)	4.7-1.3, 13 mA g <sup>-1</sup>	172	
Li <sub>2</sub> V <sub>1-x</sub> Cr <sub>x</sub> O <sub>2</sub> F (x=0.5)/C	Ball-milling	Nanoparticles	523/2e <sup>-</sup>	360(1) / 250(60)	4.7-1.3, 13 mA g <sup>-1</sup>	172	
Li <sub>2</sub> V <sub>1-x</sub> Cr <sub>x</sub> O <sub>2</sub> F (x=0.2)/Super P	Ball-milling	Nanoparticles	525/2e <sup>-</sup>	280(1) / 250(50)	4.8-1.3, 13 mA g <sup>-1</sup>	169,173	
<b>HTB</b>							
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Hydrothermal	Microparticles	225/1e <sup>-</sup>	190(1) / 136(30)	4.5-2.0, 0.1C	174	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Precipitation/IL	Nanobranchs		154(1) / 130(30)	4.5-1.6, 14 mA.g <sup>-1</sup>	175	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Precipitation/IL	Nano (10nm) /mesoporous		150(1) / 115(50)	4.5-1.6, 0.1C	176	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/SWNT	Precipitation/IL	Nanoparticles		220(1) / 143(50)	4.5-1.7, 0.1C	177	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/AB	Precipitation, ball-milling, heating	Microparticles (<1 μm)		129(1) / 102(100)	4.5-2.0, 0.1C	178	
Fe <sub>0.95</sub> Co <sub>0.05</sub> F <sub>3</sub> ·0.33H <sub>2</sub> O/AB				152(1) / 140(100)			
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/FeF <sub>3</sub> (H <sub>2</sub> O) <sub>2</sub> ·H <sub>2</sub> O/ACMB	Precipitation/heating	Spherical particles (30 μm) composite	-	179(1) / 140(50)	4.5-2.0, 23.7mA.g <sup>-1</sup>	179	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/SWNT	LIB	Precipitation/IL	225/1e <sup>-</sup>	160(1) / 140(100)	4.5-1.7, 0.1C	180	
	NIB			130(1) / 74(50)	4.0-1.2, 0.1C		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Precipitation	Nanoparticles (2-7 nm)	225/1e <sup>-</sup>	180(1) / 149(100)	4.5-2.0, 0.1C	19	
				703(1) / 166(35)	4.5-1.0, 0.1C		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/GNS	Precipitation/IL	Nanoparticles	110(1) / 113(250)	4.5-1.7, 10C	181		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Ball-milling	Nanoparticles (25-40 nm)	235(1) / 150(10)	4.5-1.8, 1C	182		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Sovothermal/Ripening/Calcination	Porous hollow microspheres	100(1) / 100(45)	4.5-1.5, 1C	183		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/rGO	Sol-gel/Microwaves	Nanoparticles (30 nm)	270(1) / 150(55)	4.5-1.5, C/10	184		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@CMK-3	Nanocasting	Nanocomposite/mesoporous	100(1) / 78(100)	4.5-1.7, 50C	185		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Solvothermal/Ti foil	Hierarchical 3D porous Microflower	126(1) / 123(50)	4.5-1.7, 3C	186		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/AB	Solvothermal	Hollow prismatic/cylindric (0.5 μm) of tiny particles (50 nm)	160(1) – 136(100)	4.5-2.0, 0.5C	187		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Precipitation	Mesoporous microsphereres	262(1) / 174(100)	4.5-2.0, 0.6C	188		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C @CNHs	Precipitation	Mesoporous nanocomposite	157(1) / 154(50)	4.5-1.7, 1C	189		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@GF-scCO <sub>2</sub>	scCO <sub>2</sub> assisted method	microflowers	675/3e <sup>-</sup>	550(1) / 145(30)	4.5-1.4, 1C	190	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@CNHs	Precipitation	3D Hierarchical nanocomposite	225/1e <sup>-</sup>	162(1) / 154(50)	4.5-1.7, 1C	189	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	Precipitation	Spindle (400x200 nm)		158(1) / 120(10)	4.5-1.8, 20 mAg <sup>-1</sup>		21
	Ball-milling	Nano (40-60 nm)		234(1) / 157(10)			
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	LIB	Solvothermal	675/3e <sup>-</sup>	276(1) / 193(50)	4.5-1.5, 1C	191	
	NIB			213(1) / 102(40)	4.0-1.0, 1C		

FeF <sub>3</sub> -0.33H <sub>2</sub> O/rGO		Solvothermal	Microparticles	225/1e <sup>-</sup>	145(1) / 100(30)	4.0-2.0, 0.1C	192
				675/3e <sup>-</sup>	700(1) / 165(30)	4.0-1.0, 0.1C	
FeF <sub>3</sub> -0.33H <sub>2</sub> O/rGO		Ball milling/reduction	Nano/composite		450(1) / 166(100)	4.3-1.3, 0.05C	193
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C		Reversed micelle method	Hierarchical mesoporous nanoflowers	225/1e <sup>-</sup>	146(1) / 143(100)	4.5-1.8, 5C	194
FeF <sub>3</sub> -0.33H <sub>2</sub> O/PGS		Precipitation	Nano (300 nm)	675/3e <sup>-</sup>	988(1) / 601(60)	4.5-1.5, 0.5C	195
Dehydrated HTB-FeF <sub>3</sub> /SWNT/C		Ionothermal fluorination	Nanodomains (4 nm) embedded in matrix		700(1) / 200(100)	4.5-1.7, 0.1C	196
HTB-FeF <sub>3</sub> /SWNT		Precipitation/IL	Nanoparticles		700(1) / 200(100)	4.5-1.3, 0.1C	196
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C (Super P)		Precipitation	nanocomposite	225/1e <sup>-</sup>	187(1) / 172(50)	4.5-2.0, 0.1C	197
FeF <sub>3</sub> -0.33H <sub>2</sub> O@C hybrids nanoreactors		Hydrothermal, heating	Nanoreactors (60 nm)	675/3e <sup>-</sup>	321(1) - 260(400)	4.5-1.5, 100 mA.g <sup>-1</sup>	198
FeF <sub>3</sub> -0.33H <sub>2</sub> O/Ag/SP		Precipitation	Microparticles	225/1e <sup>-</sup>	168(1) / 128(50)	4.5-2.0, 0.1C	199
FeF <sub>3</sub> -0.33H <sub>2</sub> O/rGO		Precipitation/heating	Nanocomposite		137(1) / 133(100)	4.5-1.8, 2C	200
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C		solvothermal	Micro/nano(30-50 nm)		146(1) / 150(50)	4.5-1.7, 0.5C	201
FeF <sub>3</sub> -0.33H <sub>2</sub> O/3D-OMC	NIB	Solvothermal	Microparticles	675/3e <sup>-</sup>	386(1) / 238(100)	4.0-1.0, 20 mA.g <sup>-1</sup>	202
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C	NIB	Solvothermal	Flowerlike mesoporous nanostructures (200 nm)	225/1e <sup>-</sup>	283(1) / 190(100)	4.5-1.5, 0.1C	203
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C		Solvothermal	Hollow nanospheres (150 nm)		87(1) / 120(100)	4.5-1.7, 0.1C	204
FeF <sub>3</sub> -0.33H <sub>2</sub> O/AIPO <sub>4</sub> /C	NIB	Solvothermal	Hollow porous microspheres	675/3e <sup>-</sup>	290(1) / 211(80)	4.0-1.2, 0.1C	205
FeF <sub>3</sub> -0.33H <sub>2</sub> O/MWCT/C	NIB	Solvothermal	Spherical nanocomposites/mesoporous		350(1) / 123(50)	4.5-1.5, 0.1C	206
FeF <sub>3</sub> -0.33H <sub>2</sub> O/MWCT/C		Precipitation	Nanocomposite (500-600 nm)		597(1) / 498(50)	4.0-1.4, 0.6C	207
FeF <sub>3</sub> -0.33H <sub>2</sub> O/CB		Solvothermal	Nanoparticles (50 nm)	225/1e <sup>-</sup>	170(1) / 130(200)	4.5-2.0, 1C	208
FeF <sub>3</sub> -0.33H <sub>2</sub> O/graphitized Carbon		Solvothermal (MOF)/thermal decomposition	Nanocomposite/porous		197(1) / 162(50)	4.5-1.7, 1C	209
FeF <sub>3</sub> -0.33H <sub>2</sub> O@3DPC		Solvothermal/heating	nanocomposite		125(1) / 101(500)	4.5-2.0, 5C	210
FeF <sub>3</sub> -0.33H <sub>2</sub> O/C		Precipitation	Microsphere/nanoplate		173(1) / 167(100)	4.5-2.0, 20 mA.g <sup>-1</sup>	211
FeF <sub>3</sub> -0.33H <sub>2</sub> O/2DrGO		Solvothermal	Nano/rGo composite		177(1) / 175(100)	4.5-1.7, 0.5C	212
FeF <sub>3</sub> -0.33H <sub>2</sub> O/CNT+graphene		Precipitation	Nano (50-100 nm)		225(1) / 193(50)	4.5-1.8, 0.2C	213
FeF <sub>3</sub> -0.33H <sub>2</sub> O/FeF <sub>3</sub> /C		Solvothermal	Hollow Yolk-like spheres	675/3e <sup>-</sup>	306(1) / 164(40)	4.2-1.5, 0.3C	214
FeF <sub>3</sub> -0.33H <sub>2</sub> O/GQDs/C		Solvothermal/Surface modification	Nanosheets	225/1e <sup>-</sup>	155(1) / 97(1000)	4.5-1.7, 2C	215
FeF <sub>3</sub> -0.33H <sub>2</sub> O/rGO	NIB	Reverse micelle method	Mesoporous/nanocomposite		227(1) / 101(00)	4.0-1.4, 1C	216
FeF <sub>3</sub> -0.33H <sub>2</sub> O/graphene/CNT		Solvothermal/IL	Micro/Nano/porous		162(1) / 120(100)	4.5-1.7, 1C	217

FeF <sub>3</sub> ·0.33H <sub>2</sub> O@3DrGO		Solvothermal/in-situ fluorination	Nanocomposite		150(1) / 100(1400)	4.5-2.0, 1.25C	218	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@NSPC		Wet impregnation/Dehydration	Micro/Porous		184(1) / 164(100)	4.5-2.0, 40 mA.g <sup>-1</sup>	219	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C		Reverse micelle method	Hierarchical nanoparticles (100 nm)	675/3e <sup>-</sup>	527(1) / 240(100)	4.2-1.5, 24 mA.g <sup>-1</sup>	220	
o-FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C (MIL-53)		Solvothermal/copyrolysis	Octahedra/Porous (1µm)	225/1e <sup>-</sup>	212(1) / 173(100)	4.5-1.5, 60 mA.g <sup>-1</sup>	221	
s-FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C (MIL-53)			Spindles/Porous (500 nm)		150(1) / 90(1000)	4.5-1.5, 474 mA.g <sup>-1</sup>		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C	LIB	Solvothermal/calcination	Raspberry-like 3D hierarchical microsized sphere	675/3e <sup>-</sup>	232(1) / 147(100)	4.5-1.5, 60 mA.g <sup>-1</sup>	222	
	NIB				125(1) / 78(1000)	4.5-1.5, 474 mA.g <sup>-1</sup>		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@CNS/SP		Precipitation/Heat treatment	Nanocomposite	225/1e <sup>-</sup>	438(1) / 284(100)	4.5-1.5, 0.1C	223	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O@CNS/LCNS/SP					675/3e <sup>-</sup>	228(1) / 150(100)		4.5-1.5, 0.1C
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C		Solvothermal (surfactants)	Micro/Nano hierarchical particles	225/1e <sup>-</sup>	175(1) / 173(200)	4.2-2.0, 1C	224	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/V <sub>2</sub> O <sub>5</sub> /C		Precipitation	Nano/composite		248(1) / 234(100)	4.2-0.01, 1C		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/MoS <sub>2</sub> /C		Precipitation	Nano/composite		213(1) / 109(100)	4.5-2.0, 1C		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/C		Precipitation	Microparticles	675/3e <sup>-</sup>	219(1) / 192(30)	4.5-2.0, 0.1C	225	
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/MoO <sub>3</sub> /C					170(1) / 141(30)	4.5-2.0, 0.1C		
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/Li <sub>3</sub> FeF <sub>6</sub> /C					Precipitation	Microparticles		496(1) / 133(50)
FeF <sub>3</sub> ·0.33H <sub>2</sub> O/TiO <sub>2</sub> /C		Solvothermal	Microparticles	424(1) / 216(50)	4.5-1.5, 0.1C			
FeF <sub>2.2</sub> O <sub>0.4</sub> □ <sub>0.4</sub>		Degradation	Microparticles	225/1e <sup>-</sup>	162(1) / 99(100)	4.5-2.0, 0.1C	228	
FeF <sub>2.66</sub> (OH) <sub>0.34</sub>		Degradation	Microparticles		503(1) / 174(100)	4.5-1.5, 0.2C		
Fe <sub>0.92</sub> Ti <sub>0.08</sub> F <sub>3</sub> ·0.33H <sub>2</sub> O/C		Precipitation	Nano (100 nm)	675/3e <sup>-</sup>	453(1) / 260(200)	4.5-1.5, 0.2C	229	
Fe <sub>0.92</sub> Mn <sub>0.08</sub> F <sub>3</sub> ·0.33H <sub>2</sub> O/C		Hydrothermal	Worm-like mesoporous structure microspheres		180(1) / 110(10)	4.2-2.0, 50 mA.g <sup>-1</sup>		230-232
Fe <sub>0.98</sub> Mn <sub>0.03</sub> F <sub>3</sub> ·0.33H <sub>2</sub> O/C		Hydrothermal	Microparticles		170(1) / 110(10)	4.0-2.0, 0.05C		
Fe <sub>0.947</sub> Ni <sub>0.08</sub> F <sub>3</sub> ·0.33H <sub>2</sub> O/C		Solvothermal	Nano/Microspheres (1µm)	225/1e <sup>-</sup>	460(1) / 295(40)	4.5-1.5, 0.1C	234,235	
Fe <sub>0.55</sub> V <sub>0.45</sub> F <sub>2.67</sub> (OH) <sub>0.33</sub> /C		Solvothermal (MW)/thermal decomposition	Microparticles		450(1) / 180(100)	4.5-1.5, 0.2C		236
<b>Pyrochlore</b>								
Fe <sub>1.9</sub> F <sub>4.75</sub> ·0.95H <sub>2</sub> O/C		Precipitation/IL, heating	Nanoparticles	221/1e <sup>-</sup>	120(1) / 90(10)	4.5-1.6, 14 mA g <sup>-1</sup>	175	
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/IL binder/graphene layer		Precipitation/IL	Nanocomposite	220/1e <sup>-</sup>	130(1) / 115(50)	4.5-1.7, 0.1C	177	
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/SWTN	LIB	Precipitation/IL	Microporous Nanoparticles	221/1e <sup>-</sup>	143(1) / 114(300)	4.5-1.7, 0.1C	240	

	NIB				250(1) / 150(50)	4.0-0.8, 0.1C	
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/SWNT(5%)		Precipitation/IL	Porous Microspheres (1 μm)	663/3e <sup>-</sup>	575(1) / 100(18)	4.5-1.0, 0.1C	241
Fe <sub>1.9</sub> F <sub>4.75</sub> ·0.95H <sub>2</sub> O/C		Microwaves/IL	Mesoporous Nanospheres (300 nm)		300(1) / 130(100)	4.5-1.6, 14 mA g <sup>-1</sup>	242
FeF <sub>3</sub> ·xH <sub>2</sub> O/AC		Precipitation, solvent exchange, heating	Porous nanospheres (<10nm)	225/1e <sup>-</sup>	222(1) / 145(50)	4.0-1.7, 50 mA g <sup>-1</sup>	243
Fe <sub>1.9</sub> F <sub>4.75</sub> ·0.95H <sub>2</sub> O/FeF <sub>3</sub> ·H <sub>2</sub> O/C		Solvothermal/IL	Mesoporous hollow Nanospheres (500 nm)	663/3e <sup>-</sup>	380(1) / 148(100)	4.5-1.6, 0.1C	244
Fe <sub>2</sub> F <sub>5</sub> ·H <sub>2</sub> O/graphene/C	NIB	Precipitation/IL	Nanospheres	221/1e <sup>-</sup>	175(1) / 145(30)	4.0-1.0, 20 mA g <sup>-1</sup>	245
Fe <sub>2</sub> F <sub>5</sub> ·H <sub>2</sub> O/graphene/C	NIB	Precipitation/IL	Nanospheres		175(1) / 145(30)	4.0-1.0, 20 mA g <sup>-1</sup>	245
FeF <sub>2.5</sub> ·0.5H <sub>2</sub> O/MWCNT		Precipitation/IL	Nanospheres (200 nm)/nanocomposite	663/3e <sup>-</sup>	325(1) / 147(100)	4.5-1.5, 40 mA g <sup>-1</sup>	246
Fe <sub>2</sub> F <sub>5</sub> ·H <sub>2</sub> O/MWCNT	NIB	Precipitation/IL	Nanospheres (20 nm)	221/1e <sup>-</sup>	127(1) / 115(50)	4.0-1.0, 100 mA g <sup>-1</sup>	247
Fe <sub>2</sub> F <sub>5</sub> ·H <sub>2</sub> O/rGO	NIB	Precipitation/IL	Nanocomposite	663/3e <sup>-</sup>	249(1) / 164(100)	4.0-1.0, 20 mA g <sup>-1</sup>	248
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/MWCNT	NIB	Precipitation/IL	Mesoporous Nanocomposite (10-100 nm)	220/1e <sup>-</sup>	194(1) / 148(100)	4.5-1.5, 0.05C	249
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/C		Solvothermal	Microspheres (1μm)	220/1e <sup>-</sup>	130(1) / 75(50)	4.0-2.0, 0.1C	250
				663/3e <sup>-</sup>	630(1) / 160(30)	3.8-1.0, 0.1C	
3D Fe <sub>2</sub> F <sub>5</sub> ·H <sub>2</sub> O@NPC		Precipitation/IL	Hybrid Nanocomposite	221/1e <sup>-</sup>	183(1) / 163(50)	4.5-1.7, 0.5C	251
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/rGO		Solvothermal	Nanospheres (5 nm) / composite	220/1e <sup>-</sup>	223(1) / 145(100)	4.5-2.0, 0.05C	252
FeF <sub>3</sub> ·0.5H <sub>2</sub> O /C		Solid gaz F <sub>2</sub> reaction	Microcrystalline powder		200(1e <sup>-</sup> ), 630(3e <sup>-</sup> )	3.5-1.0, 20 mA.g <sup>-1</sup>	158
Fe <sub>0.9</sub> Co <sub>0.1</sub> F <sub>3</sub> ·0.5H <sub>2</sub> O/C		Precipitation/IL	Nanospheres (600 nm)		227(1) / 150(200)	4.5-1.8, 0.1C	253
Fe <sub>1.95</sub> Cr <sub>0.05</sub> F <sub>5</sub> ·H <sub>2</sub> O/C	NIB	Precipitation/IL	Nano (500 nm)		357(1) / 171(100)	4.0-1.0, 0.1C	254
<b>TTB</b>							
K <sub>0.6</sub> FeF <sub>3</sub> /KB/C	NIB	Ball-milling	Nano (100 nm)	197/1e <sup>-</sup>	270(1) / 100(35)	4.5-1.5, 0.1C	255
<b>Amorphous</b>							
(3DOM)FeF <sub>3</sub> -PEDOT		Solid template method (CCT)	Hybrid nanostructure, pores (250 nm, 20-50 nm)	237/1e <sup>-</sup>	210(1) / 200(30)	4.5-2.0, 20 mA.g <sup>-1</sup>	256
FeF <sub>3</sub> ·0.5H <sub>2</sub> O/SWNT(5%)		Precipitation/IL	Nanosheets (55 nmx1-2 μm)	663/3e <sup>-</sup>	700(1) / 200(60)	4.5-1.7, 0.1C	241
FeF <sub>3</sub> /AC		Gaseous fluorination, 200°C	Hierarchized macroporous/ mesoporous texture	237/1e <sup>-</sup>	175(1) – 95(16)	4.3-2.5, 0.05C, 60°C	145

## Abbreviations

3D-OMC Three-dimensional Order Mesoporous Carbon  
3DOM-FeF<sub>3</sub> Three-dimensionally Ordered Macroporous FeF<sub>3</sub>  
3DPC 3D Porous Carbon  
AB Acetylene Black  
ACF Activated Carbon Foam  
ACMB Activated Carbon Microbead  
CB Carbon Black  
CFC Carbon Fiber Cloth  
CMK-3 Mesoporous Carbon template  
CNH Carbon Nanohorn  
CNS Carbon Nanosheets  
CNHs Hierarchical Carbon Nanohorns  
CMB Carbon MicroBubble  
CNT Carbon Nanotubes  
GC Graphitic Carbon  
GCL Graphitic Carbon Layers  
GF-scCO<sub>2</sub> Graphene Foam – Supercritical Carbon Dioxide  
GQD Graphene Quantum Dot  
IL Ionic Liquid  
KB Ketjen Black  
LCNS Pre-lithiated Carbon Nanosheets  
MHCS Mesoporous Hollow Carbon Spheres  
MWCT multi-walled carbon nanotubes  
NAN Nickel ammine nitrate NPC N doped Porous Carbon  
NSPC N, S co-doped Porous Carbon  
PAA Poly Acrylic Acid  
PEDOT Poly(3,4-ethylenedioxythiophene)  
PFPE Perfluoropolyether  
PHCNF Porous Hollow Carbon Nanofiber  
PGS Porous Carbon/Graphene Sheets  
rGO Reduced Graphene Oxide  
SPE Solid Polymer Electrolyte  
Super P  
SWNT Single-Wall Carbon Nanotubes



## **References**

- (1) Gocheva, I. D.; Kamimura, Y.; Doi, T.; Okada, S.; Yamaki, J.; Nisihda, T. Direct Synthesis of Cryolite Type  $\text{Li}_3\text{FeF}_6$  and Its Characterization as Positive Electrode in Li Cell. *Eng. Sciences Reports, Kyushu Univ.* **2009**, *31*, 7–11.
- (2) Gonzalo, E.; Kuhn, A.; García-Alvarado, F. On the Room Temperature Synthesis of Monoclinic  $\text{Li}_3\text{FeF}_6$ : A New Cathode Material for Rechargeable Lithium Batteries. *J. Power Sources* **2010**, *195*, 4990–4996. <https://doi.org/10.1016/J.JPOWSOUR.2010.02.040>.
- (3) Gonzalo, E.; Kuhn, A.; García-Alvarado, F. A Comparative Study of  $\alpha$ - and  $\beta$ - $\text{Li}_3\text{FeF}_6$ : Structure and Electrochemical Behavior. *J. Electrochem. Soc.* **2010**, *157*, A1002. <https://doi.org/10.1149/1.3454238>.
- (4) Basa, A.; Gonzalo, E.; Kuhn, A.; García-Alvarado, F. Reaching the Full Capacity of the Electrode Material  $\text{Li}_3\text{FeF}_6$  by Decreasing the Particle Size to Nanoscale. *J. Power Sources* **2012**, *197*, 260–266. <https://doi.org/10.1016/J.JPOWSOUR.2011.09.048>.
- (5) Lieser, G.; Schroeder, M.; Geßwein, H.; Winkler, V.; Glatthaar, S.; Yavuz, M.; Binder, J. R. Sol–Gel Processing and Electrochemical Characterization of Monoclinic  $\text{Li}_3\text{FeF}_6$ . *J. Sol-Gel Sci. Technol.* **2014**, *71*, 50–59. <https://doi.org/10.1007/s10971-014-3329-1>.
- (6) Shi, Y.; Sun, S.; Liu, J.; Cui, Y.; Zhuang, Q.; Chen, X. Enhanced Charge Storage of  $\text{Li}_3\text{FeF}_6$  with Carbon Nanotubes for Lithium-Ion Batteries. *RSC Adv.* **2016**, *6*, 113283–113288. <https://doi.org/10.1039/C6RA23941B>.
- (7) Basa, A.; Wojtulewski, S.; Kalska-Szostko, B.; Perkowski, M.; Gonzalo, E.; Chernyayeva, O.; Kuhn, A.; García-Alvarado, F. Carbon Coating of Air-Sensitive Insulating Transition Metal Fluorides: An Example Study on  $\alpha$ -  $\text{Li}_3\text{FeF}_6$  High-Performance Cathode for Lithium Ion Batteries. *J. Mater. Sci. Technol.* **2020**, *55*, 107–115. <https://doi.org/https://doi.org/10.1016/j.jmst.2019.10.002>.
- (8) Gocheva, I.; Chihara, K.; Okada, S.; Yamaki, J. Synthesis and Properties of  $\text{Li}_3\text{MF}_6$  ( M = V, Cr, Fe) as Positive Electrode Material for Rechargeable Batteries. *LiBD-5 Electrode Mater.* **2011**, *6*, 6–7.
- (9) Basa, A.; Gonzalo, E.; Kuhn, A.; García-Alvarado, F. Facile Synthesis of  $\beta$ - $\text{Li}_3\text{VF}_6$ : A New Electrochemically Active Lithium Insertion Material. *J. Power Sources* **2012**, *207*, 160–165. <https://doi.org/10.1016/J.JPOWSOUR.2012.01.148>.
- (10) Lieser, G.; Winkler, V.; Geßwein, H.; de Biasi, L.; Glatthaar, S.; Hoffmann, M. J.; Ehrenberg, H.; Binder, J. R. Electrochemical Characterization of Monoclinic and Orthorhombic  $\text{Li}_3\text{CrF}_6$  as Positive Electrodes in Lithium-Ion Batteries Synthesized by a Sol–Gel Process with Environmentally Benign Chemicals. *J. Power Sources* **2015**, *294*, 444–451. <https://doi.org/10.1016/J.JPOWSOUR.2015.06.036>.
- (11) Shakoob, R. A.; Lim, S. Y.; Kim, H.; Nam, K. W.; Kang, J. K.; Kang, K.; Choi, J. W. Mechanochemical Synthesis and Electrochemical Behavior of  $\text{Na}_3\text{FeF}_6$  in Sodium and Lithium

- Batteries. *Solid State Ionics* **2012**, *218*, 35–40. <https://doi.org/10.1016/j.ssi.2012.05.002>.
- (12) Sun, S.; Shi, Y.; Bian, S.; Zhuang, Q.; Liu, M.; Cui, Y. Enhanced Charge Storage of  $\text{Na}_3\text{FeF}_6$  with Carbon Nanotubes for Lithium-Ion Batteries. *Solid State Ionics* **2017**, *312*, 61–66. <https://doi.org/10.1016/J.SSI.2017.10.010>.
- (13) Yang, W.; Wang, Z.; Zhang, W.; Guo, S. Electronic-Structure Tuning of Water-Splitting Nanocatalysts. *Trends Chem.* **2019**, *1*, 259–271. <https://doi.org/10.1016/j.trechm.2019.03.006>.
- (14) Liu, W.; Wang, W.; Qin, M.; Shen, B. Successive Synthesis and Electrochemical Properties of  $\text{Na}_3\text{FeF}_6$  and  $\text{NaFeF}_3/\text{C}$  Cathode Materials for Lithium-Ion and Sodium-Ion Batteries. *Ceram. Int.* **2020**, *46*, 11436–11440. <https://doi.org/10.1016/j.ceramint.2020.01.059>.
- (15) Foley, E. E.; Wong, A.; Vincent, R. C.; Manche, A.; Zaveri, A.; Gonzalez-Correa, E.; Menard, G.; Clement, R. J. Probing Reaction Processes and Reversibility in Earth-Abundant  $\text{Na}_3\text{FeF}_6$  for Na-Ion Batteries. *Phys. Chem. Chem. Phys.* **2021**. <https://doi.org/10.1039/d1cp02763h>.
- (16) Liu, M.; Shi, Y.; Zhuang, Q. Hydrothermal Synthesis of  $\text{K}_3\text{FeF}_6$  and Its Electrochemical Characterization as Cathode Material for Lithium-Ion Batteries. *SN Appl. Sci.* **2019**, *1*, 1–9. <https://doi.org/10.1007/s42452-019-0904-7>.
- (17) Nava-Avendaño, J.; Ayllón, J. A.; Frontera, C.; Oró-Solé, J.; Estruga, M.; Molins, E.; Palacín, M. R. Low Temperature Synthesis and Characterization of Na-M-(O)-F Phases with M=Ti, V. *J. Solid State Chem.* **2015**, *226*, 286–294. <https://doi.org/10.1016/j.jssc.2015.03.006>.
- (18) Wu, C.; Li, X. X.; Wu, F.; Bai, Y.; Chen, M. Z.; Zhong, Y. Composite  $\text{FeF}_3 \cdot 3\text{H}_2\text{O}/\text{C}$  Cathode Material for Lithium Ion Battery. *Adv. Mater. Res.* **2012**, *391–392*, 1090–1094. <https://doi.org/10.4028/www.scientific.net/amr.391-392.1090>.
- (19) Liu, L.; Guo, H.; Zhou, M.; Wei, Q.; Yang, Z.; Shu, H.; Yang, X.; Tan, J.; Yan, Z.; Wang, X. A Comparison among  $\text{FeF}_3 \cdot 3\text{H}_2\text{O}$ ,  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  and  $\text{FeF}_3$  Cathode Materials for Lithium Ion Batteries: Structural, Electrochemical, and Mechanism Studies. *J. Power Sources* **2013**, *238*, 501–515. <https://doi.org/10.1016/J.JPOWSOUR.2013.04.077>.
- (20) Shi, Y. L.; Wu, N.; Shen, M. F.; Cui, Y. L.; Jiang, L.; Qiang, Y. H.; Zhuang, Q. C. Electrochemical Behavior of Iron(III) Fluoride Trihydrate as a Cathode in Lithium-Ion Batteries. *ChemElectroChem* **2014**, *1*, 645–654. <https://doi.org/10.1002/celec.201300069>.
- (21) Chen, C.; Xu, X.; Chen, S.; Zheng, B.; Shui, M.; Xu, L.; Zheng, W.; Shu, J.; Cheng, L.; Feng, L.; et al. The Preparation and Characterization of Iron Fluorides Polymorphs  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  and  $\beta\text{-FeF}_3 \cdot 3\text{H}_2\text{O}$  as Cathode Materials for Lithium-Ion Batteries. *Mater. Res. Bull.* **2015**, *64*, 187–193. <https://doi.org/10.1016/j.materresbull.2014.12.061>.
- (22) Nava-Avendaño, J.; Arroyo-De Dompablo, M. E.; Frontera, C.; Ayllón, J. A.; Palacín, M. R. Study of Sodium Manganese Fluorides as Positive Electrodes for Na-Ion Batteries. *Solid State Ionics* **2015**, *278*, 106–113. <https://doi.org/10.1016/j.ssi.2015.05.023>.
- (23) Crespi, A. M.; Somdahl, S. K.; Schmidt, C. L.; Skarstad, P. M. Evolution of Power Sources for Implantable Cardioverter Defibrillators. *J. Power Sources* **2001**, *96*, 33–38.

[https://doi.org/10.1016/S0378-7753\(01\)00499-2](https://doi.org/10.1016/S0378-7753(01)00499-2).

- (24) Sorensen, E. M.; Izumi, H. K.; Vaughey, J. T.; Stern, C. L.; Poeppelmeier, K. R.  $\text{Ag}_4\text{V}_2\text{O}_6\text{F}_2$ : An Electrochemically Active and High Silver Density Phase. *J. Am. Chem. Soc.* **2005**, *127*, 6347–6352. <https://doi.org/10.1021/ja050150f>.
- (25) Albrecht, T. A.; Sauvage, F.; Bodenez, V.; Tarascon, J. M.; Poeppelmeier, K. R. Room Temperature Synthesis of the Larger Power, High Silver Density Cathode Material  $\text{Ag}_4\text{V}_2\text{O}_6\text{F}_2$  for Implantable Cardioverter Defibrillators. *Chem. Mater.* **2009**, *21*, 3017–3020. <https://doi.org/10.1021/cm900905e>.
- (26) Donakowski, M. D.; Görne, A.; Vaughey, J. T.; Poeppelmeier, K. R.  $\text{AgNa}(\text{VO}_2\text{F}_2)_2$ : A Trioxovanadium Fluoride with Unconventional Electrochemical Properties. *J. Am. Chem. Soc.* **2013**, *135*, 9898–9906. <https://doi.org/10.1021/ja404189t>.
- (27) Dey, U. K.; Barman, N.; Ghosh, S.; Sarkar, S.; Peter, S. C.; Senguttuvan, P. Topochemical Bottom-Up Synthesis of 2D- and 3D-Sodium Iron Fluoride Frameworks. *Chem. Mater.* **2019**, *31*, 295–299. <https://doi.org/10.1021/acs.chemmater.8b04010>.
- (28) Makimura, Y.; Rougier, A.; Tarascon, J. M. Pulsed Laser Deposited Iron Fluoride Thin Films for Lithium-Ion Batteries. *Appl. Surf. Sci.* **2006**, *252*, 4587–4592. <https://doi.org/10.1016/j.apsusc.2005.06.043>.
- (29) Wang, F.; Robert, R.; Chernova, N. A.; Pereira, N.; Omenya, F.; Badway, F.; Hua, X.; Ruotolo, M.; Zhang, R.; Wu, L.; et al. Conversion Reaction Mechanisms in Lithium Ion Batteries: Study of the Binary Metal Fluoride Electrodes. *J. Am. Chem. Soc.* **2011**, *133*, 18828–18836. <https://doi.org/10.1021/ja206268a>.
- (30) Armstrong, M. J.; Panneerselvam, A.; O'Regan, C.; Morris, M. A.; Holmes, J. D. Supercritical-Fluid Synthesis of  $\text{FeF}_2$  and  $\text{CoF}_2$  Li-Ion Conversion Materials. *J. Mater. Chem. A* **2013**, *1*, 10667–10676. <https://doi.org/10.1039/c3ta12436c>.
- (31) He, K.; Zhou, Y.; Gao, P.; Wang, L.; Pereira, N.; Amatucci, G. G.; Nam, K. W.; Yang, X. Q.; Zhu, Y.; Wang, F.; et al. Sodiation via Heterogeneous Disproportionation in  $\text{FeF}_2$  Electrodes for Sodium-Ion Batteries. *ACS Nano* **2014**, *8*, 7251–7259. <https://doi.org/10.1021/nn502284y>.
- (32) Zhou, J.; Zhang, D.; Zhang, X.; Song, H.; Chen, X. Carbon-Nanotube-Encapsulated  $\text{FeF}_2$  Nanorods for High-Performance Lithium-Ion Cathode Materials. *ACS Appl. Mater. Interfaces* **2014**, *6*, 21223–21229. <https://doi.org/10.1021/am506236n>.
- (33) Zhang, Y.; Wang, L.; Li, J.; Wen, L.; He, X. A One-Pot Approach towards  $\text{FeF}_2$ -Carbon Core-Shell Composite and Its Application in Lithium Ion Batteries. *J. Alloys Compd.* **2014**, *606*, 226–230. <https://doi.org/10.1016/J.JALLCOM.2014.04.036>.
- (34) Gu, W.; Magasinski, A.; Zdyrko, B.; Yushin, G. Metal Fluorides Nanoconfined in Carbon Nanopores as Reversible High Capacity Cathodes for Li and Li-Ion Rechargeable Batteries:  $\text{FeF}_2$  as an Example. *Adv. Energy Mater.* **2015**, *5*, 1–7. <https://doi.org/10.1002/aenm.201401148>.
- (35) Song, H.; Cui, H.; Wang, C. Extremely High-Rate Capacity and Stable Cycling of a Highly

- Ordered Nanostructured Carbon-FeF<sub>2</sub> Battery Cathode. *J. Mater. Chem. A* **2015**, *3*, 22377–22384. <https://doi.org/10.1039/c5ta06297g>.
- (36) Sina, M.; Thorpe, R.; Rangan, S.; Pereira, N.; Bartynski, R. A.; Amatucci, G. G.; Cosandey, F. Investigation of SEI Layer Formation in Conversion Iron Fluoride Cathodes by Combined STEM/EELS and XPS. *J. Phys. Chem. C* **2015**, *119*, 9762–9773. <https://doi.org/10.1021/acs.jpcc.5b02058>.
- (37) Gu, W.; Borodin, O.; Zdyrko, B.; Lin, H. T.; Kim, H.; Nitta, N.; Huang, J.; Magasinski, A.; Milicev, Z.; Berdichevsky, G.; et al. Lithium-Iron Fluoride Battery with in Situ Surface Protection. *Adv. Funct. Mater.* **2016**, *26*, 1507–1516. <https://doi.org/10.1002/adfm.201504848>.
- (38) Guntlin, C. P.; Zünd, T.; Kravchyk, K. V.; Wörle, M.; Bodnarchuk, M. I.; Kovalenko, M. V. Nanocrystalline FeF<sub>3</sub> and MF<sub>2</sub> (M = Fe, Co, and Mn) from Metal Trifluoroacetates and Their Li(Na)-Ion Storage Properties. *J. Mater. Chem. A* **2017**, *5*, 7383–7393. <https://doi.org/10.1039/c7ta00862g>.
- (39) Kim, S.; Liu, J.; Sun, K.; Wang, J.; Dillon, S. J.; Braun, P. V. Improved Performance in FeF<sub>2</sub> Conversion Cathodes through Use of a Conductive 3D Scaffold and Al<sub>2</sub>O<sub>3</sub> ALD Coating. *Adv. Funct. Mater.* **2017**, *27*. <https://doi.org/10.1002/adfm.201702783>.
- (40) Ni, D.; Sun, W.; Lu, C.; Wang, Z.; Qiao, J.; Cai, H.; Liu, C.; Sun, K. Improved Rate and Cycling Performance of FeF<sub>2</sub>-RGO Hybrid Cathode with Poly (Acrylic Acid) Binder for Sodium Ion Batteries. *J. Power Sources* **2019**, *413*, 449–458. <https://doi.org/10.1016/j.jpowsour.2018.12.040>.
- (41) Huang, Q.; Turcheniuk, K.; Ren, X.; Magasinski, A.; Song, A.-Y.; Xiao, Y.; Kim, D.; Yushin, G. Cycle Stability of Conversion-Type Iron Fluoride Lithium Battery Cathode at Elevated Temperatures in Polymer Electrolyte Composites. *Nat. Mater.* **2019**, *18*, 1343–1349. <https://doi.org/10.1038/s41563-019-0472-7>.
- (42) Murugesan, V.; Cho, J. S.; Govind, N.; Andersen, A.; Olszta, M. J.; Han, K. S.; Li, G.; Lee, H.; Reed, D. M.; Sprenkle, V. L.; et al. Lithium Insertion Mechanism in Iron Fluoride Nanoparticles Prepared by Catalytic Decomposition of Fluoropolymer. *ACS Appl. Energy Mater.* **2019**, *2*, 1832–1843. <https://doi.org/10.1021/acsaem.8b01983>.
- (43) Ni, D.; Fang, L.; Sun, W.; Shi, B.; Chen, X.; Li, H.; Wang, Z.; Sun, K. FeF<sub>2</sub>@MHCS Cathodes with High Capacity and Fast Sodium Storage Based on Nanostructure Construction. *ACS Appl. Energy Mater.* **2020**, *3*, 10340–10348. <https://doi.org/10.1021/acsaem.0c00876>.
- (44) Maulana, A. Y.; Futralan, C. M.; Kim, J. MOF-Derived FeF<sub>2</sub> Nanoparticles@graphitic Carbon Undergoing in Situ Phase Transformation to FeF<sub>3</sub> as a Superior Sodium-Ion Cathode Material. *J. Alloys Compd.* **2020**, *840*, 155719. <https://doi.org/10.1016/j.jallcom.2020.155719>.
- (45) Xiao, A. W.; Lee, H. J.; Capone, I.; Robertson, A.; Wi, T. U.; Fawdon, J.; Wheeler, S.; Lee, H. W.; Grobert, N.; Pasta, M. Understanding the Conversion Mechanism and Performance of Monodisperse FeF<sub>2</sub> Nanocrystal Cathodes. *Nat. Mater.* **2020**, *19*, 644–654.

<https://doi.org/10.1038/s41563-020-0621-z>.

- (46) Li, J.; Meng, Y.; Wang, Y.; Li, X.; Lai, Y.; Guo, Y.; Wen, X.; Xiao, D. The Fluorination-Assisted Dealloying Synthesis of Porous Reduced Graphene Oxide-FeF<sub>2</sub> @carbon for High-Performance Lithium-Ion Battery and the Exploration of Its Electrochemical Mechanism. *Inorg. Chem. Front.* **2021**, c. <https://doi.org/10.1039/d1qi00273b>.
- (47) Zhang, H.; Zhou, Y. N.; Sun, Q.; Fu, Z. W. Nanostructured Nickel Fluoride Thin Film as a New Li Storage Material. *Solid State Sci.* **2008**, *10*, 1166–1172. <https://doi.org/10.1016/j.solidstatesciences.2007.12.019>.
- (48) Shi, Y. L.; Shen, M. F.; Xu, S. D.; Qiu, X. Y.; Jiang, L.; Qiang, Y. H.; Zhuang, Q. C.; Sun, S. G. Electrochemical Impedance Spectroscopic Study of the Electronic and Ionic Transport Properties of NiF<sub>2</sub>/C Composites. *Int. J. Electrochem. Sci.* **2011**, *6*, 3399–3415.
- (49) Lee, D. H.; Carroll, K. J.; Calvin, S.; Jin, S.; Meng, Y. S. Conversion Mechanism of Nickel Fluoride and NiO-Doped Nickel Fluoride in Li Ion Batteries. *Electrochim. Acta* **2012**, *59*, 213–221. <https://doi.org/10.1016/j.electacta.2011.10.105>.
- (50) Rui, K.; Wen, Z.; Lu, Y.; Jin, J.; Shen, C. One-Step Solvothermal Synthesis of Nanostructured Manganese Fluoride as an Anode for Rechargeable Lithium-Ion Batteries and Insights into the Conversion Mechanism. *Adv. Energy Mater.* **2015**, *5*, 1–11. <https://doi.org/10.1002/aenm.201401716>.
- (51) Rui, K.; Wen, Z.; Jin, J.; Huang, X. Controlled Construction of 3D Hierarchical Manganese Fluoride Nanostructures via an Oleylamine-Assisted Solvothermal Route with High Performance for Rechargeable Lithium Ion Batteries. *RSC Adv.* **2016**, *6*, 27170–27176. <https://doi.org/10.1039/c6ra03351b>.
- (52) Rui, K.; Wen, Z.; Lu, Y.; Shen, C.; Jin, J. Anchoring Nanostructured Manganese Fluoride on Few-Layer Graphene Nanosheets as Anode for Enhanced Lithium Storage. *ACS Appl. Mater. Interfaces* **2016**, *8*, 1819–1826. <https://doi.org/10.1021/acsami.5b09718>.
- (53) Bensalah, N.; Turki, D.; Kamand, F. Z.; Saoud, K. Hierarchical Nanostructured MWCNT–MnF<sub>2</sub> Composites With Stable Electrochemical Properties as Cathode Material for Lithium Ion Batteries. *Phys. Status Solidi Appl. Mater. Sci.* **2018**, *215*, 1–11. <https://doi.org/10.1002/pssa.201800151>.
- (54) Fu, Z.-W.; Li, C.-L.; Liu, W.-Y.; Ma, J.; Wang, Y.; Qin, Q.-Z. Electrochemical Reaction of Lithium with Cobalt Fluoride Thin Film Electrode. *J. Electrochem. Soc.* **2005**, *152*, E50. <https://doi.org/10.1149/1.1839512>.
- (55) Teng, Y. T.; Pramana, S. S.; Ding, J.; Wu, T.; Yazami, R. Investigation of the Conversion Mechanism of Nanosized CoF<sub>2</sub>. *Electrochim. Acta* **2013**, *107*, 301–312. <https://doi.org/10.1016/j.electacta.2013.05.107>.
- (56) Tan, J.; Liu, L.; Guo, S.; Hu, H.; Yan, Z.; Zhou, Q.; Huang, Z.; Shu, H.; Yang, X.; Wang, X. The Electrochemical Performance and Mechanism of Cobalt (II) Fluoride as Anode Material for Lithium and Sodium Ion Batteries. *Electrochim. Acta* **2015**, *168*, 225–233.

<https://doi.org/10.1016/j.electacta.2015.04.029>.

- (57) Wang, X.; Gu, W.; Lee, J. T.; Nitta, N.; Benson, J.; Magasinski, A.; Schauer, M. W.; Yushin, G. Carbon Nanotube-CoF<sub>2</sub> Multifunctional Cathode for Lithium Ion Batteries: Effect of Electrolyte on Cycle Stability. *Small* **2015**, *11*, 5164–5173. <https://doi.org/10.1002/smll.201501139>.
- (58) Guan, Q.; Cheng, J.; Li, X.; Ni, W.; Wang, B. Porous CoF<sub>2</sub> Spheres Synthesized by a One-Pot Solvothermal Method as High Capacity Cathode Materials for Lithium-Ion Batteries. *Chinese J. Chem.* **2017**, *35*, 48–54. <https://doi.org/10.1002/cjoc.201600229>.
- (59) Zhang, Q.; Huang, Y. T.; Chen, X.; Pan, A.; Cai, Z.; Liu, S.; Zhang, Y. CoF<sub>2</sub> Nanoparticles Grown on Carbon Fiber Cloth as Conversion Reaction Cathode for Lithium-Ion Batteries. *J. Alloys Compd.* **2019**, *805*, 539–544. <https://doi.org/10.1016/j.jallcom.2019.07.034>.
- (60) Cheng, Q.; Chen, Y.; Lin, X.; Liu, J.; Yuan, Z.; Cai, Y. Hybrid Cobalt(II) Fluoride Derived from a Bimetallic Zeolitic Imidazolate Framework as a High-Performance Cathode for Lithium-Ion Batteries. *J. Phys. Chem. C* **2020**, *124*, 8624–8632. <https://doi.org/10.1021/acs.jpcc.0c01292>.
- (61) Wu, F.; Srot, V.; Chen, S.; Zhang, M.; Van Aken, P. A.; Wang, Y.; Maier, J.; Yu, Y. Metal-Organic Framework-Derived Nanoconfinements of CoF<sub>2</sub> and Mixed-Conducting Wiring for High-Performance Metal Fluoride-Lithium Battery. *ACS Nano* **2021**, *15*, 1509–1518. <https://doi.org/10.1021/acsnano.0c08918>.
- (62) Seo, J. K.; Cho, H. M.; Takahara, K.; Chapman, K. W.; Borkiewicz, O. J.; Sina, M.; Shirley Meng, Y. Revisiting the Conversion Reaction Voltage and the Reversibility of the CuF<sub>2</sub> Electrode in Li-Ion Batteries. *Nano Res.* **2017**, *10*, 4232–4244. <https://doi.org/10.1007/s12274-016-1365-6>.
- (63) Krahl, T.; Marroquin Winkelmann, F.; Martin, A.; Pinna, N.; Kemnitz, E. Novel Synthesis of Anhydrous and Hydroxylated CuF<sub>2</sub> Nanoparticles and Their Potential for Lithium Ion Batteries. *Chem. Eur. J.* **2018**, *24*, 7177–7187. <https://doi.org/10.1002/chem.201800207>.
- (64) Teng, Y. T.; Wei, F.; Yazami, R. Synthesis of Ni<sub>x</sub>Co<sub>(1-x)</sub>F<sub>2</sub> (x = 0, 0.25, 0.50, 0.75, 1.0) and Application in Lithium Ion Batteries. *J. Alloys Compd.* **2015**, *653*, 434–443. <https://doi.org/10.1016/j.jallcom.2015.09.036>.
- (65) Huang, Q.; Pollard, T. P.; Ren, X.; Kim, D.; Magasinski, A.; Borodin, O.; Yushin, G. Fading Mechanisms and Voltage Hysteresis in FeF<sub>2</sub>-NiF<sub>2</sub> Solid Solution Cathodes for Lithium and Lithium-Ion Batteries. *Small* **2019**, *15*, 1–11. <https://doi.org/10.1002/smll.201804670>.
- (66) Wang, F.; Kim, S. W.; Seo, D. H.; Kang, K.; Wang, L.; Su, D.; Vajo, J. J.; Wang, J.; Graetz, J. Ternary Metal Fluorides as High-Energy Cathodes with Low Cycling Hysteresis. *Nat. Commun.* **2015**, *6*, 1–9. <https://doi.org/10.1038/ncomms7668>.
- (67) Villa, C.; Kim, S.; Lu, Y.; Dravid, V. P.; Wu, J. Cu-Substituted NiF<sub>2</sub> as a Cathode Material for Li-Ion Batteries. *ACS Appl. Mater. Interfaces* **2019**, *11*, 647–654. <https://doi.org/10.1021/acsmi.8b15791>.
- (68) Badway, F.; Mansour, A. N.; Pereira, N.; Al-Sharab, J. F.; Cosandey, F.; Plitz, I.; Amatucci,

- G. G. Structure and Electrochemistry of Copper Fluoride Nanocomposites Utilizing Mixed Conducting Matrices. *Chem. Mater.* **2007**, *19*, 4129–4141. <https://doi.org/10.1021/cm070421g>.
- (69) Pereira, N.; Badway, F.; Wartelsky, M.; Gunn, S.; Amatucci, G. G. Iron Oxyfluorides as High Capacity Cathode Materials for Lithium Batteries. *J. Electrochem. Soc.* **2009**, *156*, A407–A416. <https://doi.org/10.1149/1.3106132>.
- (70) Kitajou, A.; Komatsu, H.; Nagano, R.; Okada, S. Synthesis of FeOF Using Roll-Quenching Method and the Cathode Properties for Lithium-Ion Battery. *J. Power Sources* **2013**, *243*, 494–498. <https://doi.org/10.1016/j.jpowsour.2013.06.053>.
- (71) Kim, S. W.; Pereira, N.; Chernova, N. A.; Omenya, F.; Gao, P.; Whittingham, M. S.; Amatucci, G. G.; Su, D.; Wang, F. Structure Stabilization by Mixed Anions in Oxyfluoride Cathodes for High-Energy Lithium Batteries. *ACS Nano* **2015**, *9*, 10076–10084. <https://doi.org/10.1021/acs.nano.5b03643>.
- (72) Fan, X.; Luo, C.; Lamb, J.; Zhu, Y.; Xu, K.; Wang, C. PEDOT Encapsulated FeOF Nanorod Cathodes for High Energy Lithium-Ion Batteries. *Nano Lett.* **2015**, *15*, 7650–7656. <https://doi.org/10.1021/acs.nanolett.5b03601>.
- (73) Wang, L.-P.; Wang, T.-S.; Zhang, X.-D.; Liang, J.-Y.; Jiang, L.; Yin, Y.-X.; Guo, Y.-G.; Wang, C.-R. Iron Oxyfluorides as Lithium-Free Cathode Materials for Solid-State Li Metal Batteries. *J. Mater. Chem. A* **2017**, *5*, 18464–18468. <https://doi.org/10.1039/C7TA05138G>.
- (74) Li, W.; Li, Y.; Fang, M.; Yao, X.; Li, T.; Shui, M.; Shu, J. The Facile in Situ Preparation and Characterization of C/FeOF/FeF<sub>3</sub> Nanocomposites as LIB Cathode Materials. *Ionics (Kiel)*. **2018**, *24*, 1561–1569. <https://doi.org/10.1007/s11581-017-2334-0>.
- (75) Zhai, J.; Lei, Z.; Sun, K. 3D Starfish-like FeOF on Graphene Sheets: Engineered Synthesis and Lithium Storage Performance. *Chem. Eur. J.* **2019**, *25*, 1–8. <https://doi.org/10.1002/chem.201900948>.
- (76) Jung, S. K.; Hwang, I.; Cho, S. P.; Oh, K.; Ku, K.; Choi, I. R.; Kang, K. New Iron-Based Intercalation Host for Lithium-Ion Batteries. *Chem. Mater.* **2018**, *30*, 1956–1964. <https://doi.org/10.1021/acs.chemmater.7b05017>.
- (77) Zhou, Y. N.; Sina, M.; Pereira, N.; Yu, X.; Amatucci, G. G.; Yang, X. Q.; Cosandey, F.; Nam, K. W. FeO<sub>0.7</sub>F<sub>1.3</sub>/C Nanocomposite as a High-Capacity Cathode Material for Sodium-Ion Batteries. *Adv. Funct. Mater.* **2015**, *25*, 696–703. <https://doi.org/10.1002/adfm.201403241>.
- (78) Zhu, J.; Deng, D. Wet-Chemical Synthesis of Phase-Pure FeOF Nanorods as High-Capacity Cathodes for Sodium-Ion Batteries. *Angew. Chemie - Int. Ed.* **2015**, *54*, 3079–3083. <https://doi.org/10.1002/anie.201410572>.
- (79) Fu, S. Y.; Li, Y. Z.; Chu, W.; Yang, Y. M.; Tong, D. G.; Le Zeng, Q. Mesoporous Amorphous FeOF Nanococoons for High-Rate and Long-Life Rechargeable Sodium-Ion Batteries. *J. Mater. Chem. A* **2015**, *3*, 16716–16727. <https://doi.org/10.1039/c5ta04288g>.
- (80) Kitajou, A.; Zhao, L.; Nagano, R.; Inoishi, A.; Kobayashi, E.; Okada, S. Electrochemical

Performance and Thermal Stability of Iron Oxyfluoride (FeOF) for Sodium-Ion Batteries. *Batteries* **2018**, *4*, 68. <https://doi.org/10.3390/batteries4040068>.

- (81) Park, M.; Shim, J. H.; Kim, H.; Park, H.; Kim, N.; Kim, J. FeOF Ellipsoidal Nanoparticles Anchored on Reduced Graphene Oxides as a Cathode Material for Sodium-Ion Batteries. *J. Power Sources* **2018**, *396*, 551–558. <https://doi.org/10.1016/j.jpowsour.2018.06.071>.
- (82) Maulana, A. Y.; Kim, J. FeOF Nanoparticles Wrapped in Graphitic Carbon Layers in Situ Prepared from Fe-MIL-88B as a Cathode Material for Sodium-Ion Batteries. *Carbon N. Y.* **2019**, *149*, 483–491. <https://doi.org/10.1016/j.carbon.2019.04.081>.
- (83) Maulana, A. Y.; Song, J.; Lee, D. W.; Lee, C. E.; Kim, J. Enhanced Electrochemical Performance of Graphitic Carbon-Wrapped Spherical FeOF Nanoparticles Using Maleopimaric Acid as a Cathode Material for Sodium-Ion Batteries. *J. Mater. Sci. Technol.* **2021**, *85*, 184–193. <https://doi.org/10.1016/j.jmst.2021.01.023>.
- (84) Fan, X.; Hu, E.; Ji, X.; Zhu, Y.; Han, F.; Hwang, S.; Liu, J.; Bak, S.; Ma, Z.; Gao, T.; et al. High Energy-Density and Reversibility of Iron Fluoride Cathode Enabled via an Intercalation-Extrusion Reaction. *Nat. Commun.* **2018**, *9*, 2324. <https://doi.org/10.1038/s41467-018-04476-2>.
- (85) Twu, N.; Li, X.; Moore, C.; Ceder, G. Synthesis and Lithiation Mechanisms of Dirutile and Rutile  $\text{LiMnF}_4$ : Two New Conversion Cathode Materials. *J. Electrochem. Soc.* **2013**, *160*, A1944–A1951. <https://doi.org/10.1149/2.022311jes>.
- (86) Liao, P.; Li, J.; Dahn, J. R. Lithium Intercalation in  $\text{LiFe}_2\text{F}_6$  and  $\text{LiMgFeF}_6$  Disordered Trirutile-Type Phases. *J. Electrochem. Soc.* **2010**, *157*, a355–a361. <https://doi.org/10.1149/1.3294788>.
- (87) Lieser, G.; Dräger, C.; De Biasi, L.; Indris, S.; Geßwein, H.; Glatthaar, S.; Hoffmann, M. J.; Ehrenberg, H.; Binder, J. R. Direct Synthesis of Trirutile-Type  $\text{LiMgFeF}_6$  and Its Electrochemical Characterization as Positive Electrode in Lithium-Ion Batteries. *J. Power Sources* **2015**, *274*, 1200–1207. <https://doi.org/10.1016/j.jpowsour.2014.10.151>.
- (88) Lieser, G.; Dräger, C.; Schroeder, M.; Indris, S.; de Biasi, L.; Geßwein, H.; Glatthaar, S.; Ehrenberg, H.; Binder, J. R. Sol-Gel Based Synthesis of  $\text{LiNiFeF}_6$  and Its Electrochemical Characterization. *J. Electrochem. Soc.* **2014**, *161*, A1071–A1077. <https://doi.org/10.1149/2.070406jes>.
- (89) Gocheva, I. D.; Doi, T.; Okada, S.; Yamaki, J. ichi. Electrochemical Properties of Trirutile-Type  $\text{Li}_2\text{TiF}_6$  as Cathode Active Material in Li-Ion Batteries. *Electrochemistry* **2010**, *78*, 471–474.
- (90) De Biasi, L.; Lieser, G.; Dräger, C.; Indris, S.; Rana, J.; Schumacher, G.; Mönig, R.; Ehrenberg, H.; Binder, J. R.; Geßwein, H.  $\text{LiCaFeF}_6$ : A Zero-Strain Cathode Material for Use in Li-Ion Batteries. *J. Power Sources* **2017**, *362*, 192–201. <https://doi.org/10.1016/j.jpowsour.2017.07.007>.
- (91) a- Lieser, G.; De Biasi, L.; Geßwein, H.; Indris, S.; Dräger, C.; Schroeder, M.; Glatthaar, S.; Ehrenberg, H.; Binder, J. R. Electrochemical Characterization of  $\text{LiMnFeF}_6$  for Use as Positive



Electrode in Lithium-Ion Batteries. *J. Electrochem. Soc.* **2014**, *161*, A1869–A1876. <https://doi.org/10.1149/2.0651412jes>.

b- Lemoine, K.; Nagatani, Y.; Grenèche, J.-M.; Inaguma, Y. High-Pressure Synthesis of Trigonal LiFe<sub>2</sub>F<sub>6</sub>: New Iron Fluoride with Li<sup>+</sup> Tunnels as a Potential Cathode for Lithium-Ion Batteries. *J. Phys. Chem. C* **2022**, *126*, 8248–8255. <https://doi.org/10.1021/acs.jpcc.2c01516>.

- (92) Lieser, G.; de Biasi, L.; Scheuermann, M.; Winkler, V.; Eisenhardt, S.; Glatthaar, S.; Indris, S.; Geßwein, H.; Hoffmann, M. J.; Ehrenberg, H.; et al. Sol-Gel Processing and Electrochemical Conversion of Inverse Spinel-Type Li<sub>2</sub>NiF<sub>4</sub>. *J. Electrochem. Soc.* **2015**, *162*, A679–A686. <https://doi.org/10.1149/2.0591504jes>.
- (93) Zhang, M. Fabrication of Li<sub>2</sub>NiF<sub>4</sub>-PEDOT Nanocomposites as Conversion Cathodes for Lithium-Ion Batteries. *J. Alloys Compd.* **2017**, *723*, 139–145. <https://doi.org/10.1016/j.jallcom.2017.06.212>.
- (94) Kimura, N.; Nasu, H.; Kohno, Y.; Tomita, Y. Synthesis and Charge-Discharge Properties of 2LiF-NiF<sub>2</sub> Composite and Li<sub>2</sub>NiF<sub>4</sub> as a Cathode Material for Li-Ion Batteries. *Trans. Mater. Res. Soc. Japan* **2020**, *45*, 15–18. <https://doi.org/10.14723/tmrsj.45.15>.
- (95) Zhang, L.; Chen, G.; Berg, E. J.; Tarascon, J. M. Triggering the In Situ Electrochemical Formation of High Capacity Cathode Material from MnO. *Adv. Energy Mater.* **2017**, *7*, 1–6. <https://doi.org/10.1002/aenm.201602200>.
- (96) Zhang, L.; Dambournet, D.; Iadecola, A.; Batuk, D.; Borkiewicz, O. J.; Wiaderek, K. M.; Salager, E.; Shao, M.; Chen, G.; Tarascon, J. M. Origin of the High Capacity Manganese-Based Oxyfluoride Electrodes for Rechargeable Batteries. *Chem. Mater.* **2018**, *30*, 5362–5372. <https://doi.org/10.1021/acs.chemmater.8b02182>.
- (97) Bervas, M.; Badway, F.; Klein, L. C.; Amatucci, G. G. Bismuth Fluoride Nanocomposite as a Positive Electrode Material for Rechargeable Lithium Batteries. *Electrochem. Solid-State Lett.* **2005**, *8*, A179. <https://doi.org/10.1149/1.1861040>.
- (98) Bervas, M.; Mansour, A. N.; Yoon, W.-S.; Al-Sharab, J. F.; Badway, F.; Cosandey, F.; Klein, L. C.; Amatucci, G. G. Investigation of the Lithiation and Delithiation Conversion Mechanisms of Bismuth Fluoride Nanocomposites. *J. Electrochem. Soc.* **2006**, *153*, A799. <https://doi.org/10.1149/1.2167951>.
- (99) Konishi, H.; Minato, T.; Abe, T.; Ogumi, Z. Difference of Rate Performance between Discharge and Charge Reactions for Bismuth Fluoride Electrode in Lithium-Ion Battery. *J. Electroanal. Chem.* **2017**, *806*, 82–87. <https://doi.org/10.1016/j.jelechem.2017.10.051>.
- (100) Bervas, M.; Klein, L. C.; Amatucci, G. G. Reversible Conversion Reactions with Lithium in Bismuth Oxyfluoride Nanocomposites. *J. Electrochem. Soc.* **2005**, *153*, A159. <https://doi.org/10.1149/1.2133712>.
- (101) Ni, D.; Sun, W.; Xie, L.; Fan, Q.; Wang, Z.; Sun, K. Bismuth Oxyfluoride @ CMK-3 Nanocomposite as Cathode for Lithium Ion Batteries. *J. Power Sources* **2018**, *374*, 166–174.

<https://doi.org/10.1016/j.jpowsour.2017.11.017>.

- (102) Park, H.; Lee, Y.; Cho, M. K.; Kang, J.; Ko, W.; Jung, Y. H.; Jeon, T. Y.; Hong, J.; Kim, H.; Myung, S. T.; et al. Na<sub>2</sub>Fe<sub>2</sub>F<sub>7</sub>: A Fluoride-Based Cathode for High Power and Long Life Na-Ion Batteries. *Energy Environ. Sci.* **2021**, *14*, 1469–1479. <https://doi.org/10.1039/d0ee02803g>.
- (103) Liao, J.; Han, J.; Xu, J.; Du, Y.; Sun, Y.; Duan, L.; Zhou, X. Scalable Synthesis of Na<sub>2</sub>MVF<sub>7</sub> (M = Mn, Fe, and Co) as High-Performance Cathode Materials for Sodium-Ion Batteries. *Chem. Commun.* **2021**, *57*, 11497–11500. <https://doi.org/10.1039/d1cc04449d>.
- (104) Gocheva, I. D.; Nishijima, M.; Doi, T.; Okada, S.; Yamaki, J. ichi; Nishida, T. Mechanochemical Synthesis of NaMF<sub>3</sub> (M = Fe, Mn, Ni) and Their Electrochemical Properties as Positive Electrode Materials for Sodium Batteries. *J. Power Sources* **2009**, *187*, 247–252. <https://doi.org/10.1016/j.jpowsour.2008.10.110>.
- (105) Yamada, Y.; Doi, T.; Tanaka, I.; Okada, S.; Yamaki, J. Liquid-Phase Synthesis of Highly Dispersed NaFeF<sub>3</sub> Particles and Their Electrochemical Properties for Sodium-Ion Batteries. *J. Power Sources* **2011**, *196*, 4837–4841. <https://doi.org/10.1016/J.JPOWSOUR.2011.01.060>.
- (106) Kitajou, A.; Komatsu, H.; Chihara, K.; Gocheva, I. D.; Okada, S.; Yamaki, J. Novel Synthesis and Electrochemical Properties of Perovskite-Type NaFeF<sub>3</sub> for a Sodium-Ion Battery. *J. Power Sources* **2012**, *198*, 389–392. <https://doi.org/10.1016/J.JPOWSOUR.2011.09.064>.
- (107) Dimov, N.; Nishimura, A.; Chihara, K.; Kitajou, A.; Gocheva, I. D.; Okada, S. Transition Metal NaMF<sub>3</sub> Compounds as Model Systems for Studying the Feasibility of Ternary Li-M-F and Na-M-F Single Phases as Cathodes for Lithium-Ion and Sodium-Ion Batteries. *Electrochim. Acta* **2013**, *110*, 214–220. <https://doi.org/10.1016/J.ELECTACTA.2013.05.103>.
- (108) Kitajou, A.; Ishado, Y.; Yamashita, T.; Momida, H.; Oguchi, T.; Okada, S. Cathode Properties of Perovskite-Type NaMF<sub>3</sub> (M = Fe, Mn, and Co) Prepared by Mechanical Ball Milling for Sodium-Ion Battery. *Electrochim. Acta* **2017**, *245*, 424–429. <https://doi.org/10.1016/J.ELECTACTA.2017.05.153>.
- (109) Martin, A.; Doublet, M. L.; Kemnitz, E.; Pinna, N. Reversible Sodium and Lithium Insertion in Iron Fluoride Perovskites. *Adv. Funct. Mater.* **2018**, *28*, 1802057. <https://doi.org/10.1002/adfm.201802057>.
- (110) Kravchyk, K. V.; Zünd, T.; Wörle, M.; Kovalenko, M. V.; Bodnarchuk, M. I. NaFeF<sub>3</sub> Nanoplates as Low-Cost Sodium and Lithium Cathode Materials for Stationary Energy Storage. *Chem. Mater.* **2018**, *30*, 1825–1829. <https://doi.org/10.1021/acs.chemmater.7b04743>.
- (111) Oh, J.; Jang, J.; Lim, E.; Jo, C.; Chun, jinyoung. Synthesis of Sodium Cobalt Fluoride/Reduced Graphene Oxide (NaCoF<sub>3</sub>/RGO) Nanocomposites and Investigation of Their Electrochemical Properties as Cathodes for Li-Ion Batteries. *Materials (Basel)*. **2021**, *14*, 547. <https://doi.org/10.3390/ma14030547>.
- (112) Cao, D.; Yin, C.; Shi, D.; Fu, Z.; Zhang, J.; Li, C. Cubic Perovskite Fluoride as Open

- Framework Cathode for Na-Ion Batteries. *Adv. Funct. Mater.* **2017**, *27*, 1–9. <https://doi.org/10.1002/adfm.201701130>.
- (113) Yi, T.; Chen, W.; Cheng, L.; Bayliss, R. D.; Lin, F.; Plews, M. R.; Nordlund, D.; Doeff, M. M.; Persson, K. A.; Cabana, J. Investigating the Intercalation Chemistry of Alkali Ions in Fluoride Perovskites. *Chem. Mater.* **2017**, *29*, 1561–1568. <https://doi.org/10.1021/acs.chemmater.6b04181>.
- (114) Tong, W.; Amatucci, G. G. Silver Copper Fluoride: A Novel Perovskite Cathode for Lithium Batteries. *J. Power Sources* **2017**, *362*, 86–91. <https://doi.org/10.1016/j.jpowsour.2017.07.008>.
- (115) Badway, F.; Pereira, N.; Cosandey, F.; Amatucci, G. G. Carbon-Metal Fluoride Nanocomposites: Structure and Electrochemistry of  $\text{FeF}_3$ :C. *J. Electrochem. Soc.* **2003**, *150*, A1209. <https://doi.org/10.1149/1.1596162>.
- (116) Badway, F.; Cosandey, F.; Pereira, N.; Amatucci, G. G. Carbon Metal Fluoride Nanocomposites: High-Capacity Reversible Metal Fluoride Conversion Materials as Rechargeable Positive Electrodes for Li Batteries. *J. Electrochem. Soc.* **2003**, *150*, A1209–A1218. <https://doi.org/10.1149/1.1602454>.
- (117) Plitz, I.; Badway, F.; Al-Sharab, J.; DuPasquier, A.; Cosandey, F.; Amatucci, G. G. Structure and Electrochemistry of Carbon-Metal Fluoride Nanocomposites Fabricated by Solid-State Redox Conversion Reaction. *J. Electrochem. Soc.* **2005**, *152*, A307. <https://doi.org/10.1149/1.1842035>.
- (118) Nishijima, M.; Gocheva, I. D.; Okada, S.; Doi, T.; Yamaki, J. ichi; Nishida, T. Cathode Properties of Metal Trifluorides in Li and Na Secondary Batteries. *J. Power Sources* **2009**, *190*, 558–562. <https://doi.org/10.1016/j.jpowsour.2009.01.051>.
- (119) Kim, S. W.; Seo, D. H.; Gwon, H.; Kim, J.; Kang, K. Fabrication of  $\text{FeF}_3$  Nanoflowers on CNT Branches and Their Application to High Power Lithium Rechargeable Batteries. *Adv. Mater.* **2010**, *22*, 5260–5264. <https://doi.org/10.1002/adma.201002879>.
- (120) Li, T.; Li, L.; Cao, Y. L.; Ai, X. P.; Yang, H. X. Reversible Three-Electron Redox Behaviors of  $\text{FeF}_3$  Nanocrystals as High-Capacity Cathode-Active Materials for Li-Ion Batteries. *J. Phys. Chem. C* **2010**, *114*, 3190–3195. <https://doi.org/10.1021/jp908741d>.
- (121) Yabuuchi, N.; Sugano, M.; Yamakawa, Y.; Nakai, I.; Sakamoto, K.; Muramatsu, H.; Komaba, S. Effect of Heat-Treatment Process on  $\text{FeF}_3$  Nanocomposite Electrodes for Rechargeable Li Batteries. *J. Mater. Chem.* **2011**, *21*, 10035–10041. <https://doi.org/10.1039/c0jm04231e>.
- (122) Li, L.; Meng, F.; Jin, S. High-Capacity Lithium-Ion Battery Conversion Cathodes Based on Iron Fluoride Nanowires and Insights into the Conversion Mechanism. *Nano Lett.* **2012**, *12*, 6030–6037. <https://doi.org/10.1021/nl303630p>.
- (123) Shi, Y. L.; Shen, M. F.; Xu, S. D.; Zhuang, Q. C.; Jiang, L.; Qiang, Y. H. Electrochemical Impedance Spectroscopy Investigation of the  $\text{FeF}_3$ /C Cathode for Lithium-Ion Batteries. *Solid State Ionics* **2012**, *222–223*, 23–30. <https://doi.org/10.1016/j.ssi.2012.06.024>.

- (124) Zhang, W.; Ma, L.; Yue, H.; Yang, Y. Synthesis and Characterization of in Situ Fe<sub>2</sub>O<sub>3</sub>-Coated FeF<sub>3</sub> Cathode Materials for Rechargeable Lithium Batteries. *J. Mater. Chem.* **2012**, *22*, 24769–24775. <https://doi.org/10.1039/c2jm34391f>.
- (125) Torii, H.; Uematsu, K.; Ishigaki, T.; Toda, K.; Sato, M. Synthesis of FeF<sub>3</sub> Fluoride Electrode Material Using Polytetrafluoroethylene. *J. Ceram. Soc. Japan* **2014**, *122*, 473–476. <https://doi.org/10.2109/jcersj2.122.473>.
- (126) Myung, S. T.; Sakurada, S.; Yashiro, H.; Sun, Y. K. Iron Trifluoride Synthesized via Evaporation Method and Its Application to Rechargeable Lithium Batteries. *J. Power Sources* **2013**, *223*, 1–8. <https://doi.org/10.1016/j.jpowsour.2012.09.027>.
- (127) Ma, R.; Lu, Z.; Wang, C.; Wang, H. E.; Yang, S.; Xi, L.; Chung, J. C. Y. Large-Scale Fabrication of Graphene-Wrapped FeF<sub>3</sub> Nanocrystals as Cathode Materials for Lithium Ion Batteries. *Nanoscale* **2013**, *5*, 6338–6343. <https://doi.org/10.1039/c3nr00380a>.
- (128) Ma, R.; Wang, M.; Tao, P.; Wang, Y.; Cao, C.; Shan, G.; Yang, S.; Xi, L.; Chung, J. C. Y.; Lu, Z. Fabrication of FeF<sub>3</sub> Nanocrystals Dispersed into a Porous Carbon Matrix as a High Performance Cathode Material for Lithium Ion Batteries. *J. Mater. Chem. A* **2013**, *1*, 15060–15067. <https://doi.org/10.1039/c3ta13086j>.
- (129) Chu, Q.; Xing, Z.; Ren, X.; Asiri, A. M.; Al-Youbi, A. O.; Alamry, K. A.; Sun, X. Reduced Graphene Oxide Decorated with FeF<sub>3</sub> Nanoparticles: Facile Synthesis and Application as a High Capacity Cathode Material for Rechargeable Lithium Batteries. *Electrochim. Acta* **2013**, *111*, 80–85. <https://doi.org/10.1016/j.electacta.2013.08.006>.
- (130) Liu, J.; Wan, Y.; Liu, W.; Ma, Z.; Ji, S.; Wang, J.; Zhou, Y.; Hodgson, P.; Li, Y. Mild and Cost-Effective Synthesis of Iron Fluoride-Graphene Nanocomposites for High-Rate Li-Ion Battery Cathodes. *J. Mater. Chem. A* **2013**, *1*, 1969–1975. <https://doi.org/10.1039/c2ta00823h>.
- (131) Jung, H.; Shin, J.; Chae, C.; Lee, J. K.; Kim, J. FeF<sub>3</sub>/Ordered Mesoporous Carbon (OMC) Nanocomposites for Lithium Ion Batteries with Enhanced Electrochemical Performance. *J. Phys. Chem. C* **2013**, *117*, 14939–14946. <https://doi.org/10.1021/jp4023162>.
- (132) Ma, D. long; Wang, H. guo; Li, Y.; Xu, D.; Yuan, S.; Huang, X. lei; Zhang, X. bo; Zhang, Y. In Situ Generated FeF<sub>3</sub> in Homogeneous Iron Matrix toward High-Performance Cathode Material for Sodium-Ion Batteries. *Nano Energy* **2014**, *10*, 295–304. <https://doi.org/10.1016/j.nanoen.2014.10.004>.
- (133) Bao, T.; Zhong, H.; Zheng, H.; Zhan, H.; Zhou, Y. One-Pot Synthesis of FeF<sub>3</sub> Graphene Composite for Sodium Secondary Batteries. *Mater. Lett.* **2015**, *158*, 21–24. <https://doi.org/10.1016/j.matlet.2015.05.041>.
- (134) Bao, T.; Zhong, H.; Zheng, H.; Zhan, H.; Zhou, Y. In-Situ Synthesis of FeF<sub>3</sub>/Graphene Composite for High-Rate Lithium Secondary Batteries. *Electrochim. Acta* **2015**, *176*, 215–221. <https://doi.org/10.1016/j.electacta.2015.06.125>.
- (135) Song, H.; Yang, G.; Cui, H.; Wang, C. Honeycomb-like Porous Iron Fluoride Hybrid Nanostructures: Excellent Li-Storage Properties and Investigation of the Multi-Electron

Reversible Conversion Reaction Mechanism. *J. Mater. Chem. A* **2015**, *3*, 19832–19841. <https://doi.org/10.1039/c5ta04900h>.

- (136) Shen, Y.; Wang, X.; Hu, H.; Jiang, M.; Bai, Y.; Yang, X.; Shu, H. Sheet-like Structure FeF<sub>3</sub>/Graphene Composite as Novel Cathode Material for Na Ion Batteries. *RSC Adv.* **2015**, *5*, 38277–38282. <https://doi.org/10.1039/c5ra02235e>.
- (137) Jung, H.; Song, H.; Kim, T.; Lee, J. K.; Kim, J. FeF<sub>3</sub> Microspheres Anchored on Reduced Graphene Oxide as a High Performance Cathode Material for Lithium Ion Batteries. *J. Alloys Compd.* **2015**, *647*, 750–755. <https://doi.org/10.1016/j.jallcom.2015.06.191>.
- (138) Kim, Y. K.; Lee, J. K.; Kim, J. FeF<sub>3</sub> Nanoparticles Embedded in Activated Carbon Foam (ACF) as a Cathode Material with Enhanced Electrochemical Performance for Lithium Ion Batteries. *Bull. Korean Chem. Soc.* **2015**, *36*, 1878–1884. <https://doi.org/10.1002/bkcs.10366>.
- (139) Jiang, J.; Li, L.; Xu, M.; Zhu, J.; Li, C. M. FeF<sub>3</sub>@Thin Nickel Ammine Nitrate Matrix: Smart Configurations and Applications as Superior Cathodes for Li-Ion Batteries. *ACS Appl. Mater. Interfaces* **2016**, *8*, 16240–16247. <https://doi.org/10.1021/acsami.6b03949>.
- (140) Kim, T.; Jae, W. J.; Kim, H.; Park, M.; Han, J. M.; Kim, J. A Cathode Material for Lithium-Ion Batteries Based on Graphitized Carbon-Wrapped FeF<sub>3</sub> Nanoparticles Prepared by Facile Polymerization. *J. Mater. Chem. A* **2016**, *4*, 14857–14864. <https://doi.org/10.1039/c6ta06696h>.
- (141) Fan, X.; Zhu, Y.; Luo, C.; Gao, T.; Suo, L.; Liou, S. C.; Xu, K.; Wang, C. In Situ Lithiated FeF<sub>3</sub>/C Nanocomposite as High Energy Conversion-Reaction Cathode for Lithium-Ion Batteries. *J. Power Sources* **2016**, *307*, 435–442. <https://doi.org/10.1016/j.jpowsour.2016.01.004>.
- (142) Tawa, S.; Yamamoto, T.; Matsumoto, K.; Hagiwara, R. Iron(III) Fluoride Synthesized by a Fluorolysis Method and Its Electrochemical Properties as a Positive Electrode Material for Lithium Secondary Batteries. *J. Fluor. Chem.* **2016**, *184*, 75–81. <https://doi.org/10.1016/j.jfluchem.2016.02.009>.
- (143) Lee, J.; Kang, B. Novel and Scalable Solid-State Synthesis of a Nanocrystalline FeF<sub>3</sub>/C Composite and Its Excellent Electrochemical Performance. *Chem. Commun.* **2016**, *52*, 9414–9417. <https://doi.org/10.1039/c6cc03706b>.
- (144) Kumagae, K.; Okazaki, K.; Matsui, K.; Horino, H.; Hirai, T.; Yamaki, J.; Ogumi, Z. Improvement of Cycling Performance of FeF<sub>3</sub>-Based Lithium-Ion Battery by Boron-Based Additives. *J. Electrochem. Soc.* **2016**, *163*, A1633–A1636. <https://doi.org/10.1149/2.0871608jes>.
- (145) Guérin, K.; Delbègue, D.; Louvain, N.; Doubtsof, L.; Hamwi, A.; Laik, B.; Pereira-Ramos, J. P.; Tahar-Sougrati, M.; Jumas, J. C.; Willmann, P.; et al. Rhombohedral Iron Trifluoride with a Hierarchized Macroporous/Mesoporous Texture from Gaseous Fluorination of Iron Disilicide. *Mater. Chem. Phys.* **2016**, *173*, 355–363. <https://doi.org/10.1016/j.matchemphys.2016.02.023>.
- (146) Li, Y.; Yao, F.; Cao, Y.; Yang, H.; Feng, Y.; Feng, W. The Mediated Synthesis of FeF<sub>3</sub>

Nanocrystals through  $(\text{NH}_4)_3\text{FeF}_6$  Precursors as the Cathode Material for High Power Lithium Ion Batteries. *Electrochim. Acta* **2017**, *253*, 545–553. <https://doi.org/10.1016/j.electacta.2017.09.081>.

- (147) Zhou, X.; Sun, H.; Zhou, H.; Xu, Z.; Yang, J. Enhancing Cycling Performance of  $\text{FeF}_3$  Cathode by Introducing a Lightweight High Conductive Adsorbable Interlayer. *J. Alloys Compd.* **2017**, *723*, 317–326. <https://doi.org/10.1016/j.jallcom.2017.06.266>.
- (148) Kitajou, A.; Tanaka, I.; Tanaka, Y.; Kobayashi, E.; Setoyama, H.; Okajima, T.; Okada, S. Discharge and Charge Reaction of Perovskite-Type  $\text{MF}_3$  (M = Fe and Ti) Cathodes for Lithium-Ion Batteries. *Electrochemistry* **2017**, *85*, 472–477. <https://doi.org/10.5796/electrochemistry.85.472>.
- (149) Sun, H.; Zhou, H.; Xu, Z.; Ding, J.; Yang, J.; Zhou, X. Preparation of Anhydrous Iron Fluoride with Porous Fusiform Structure and Its Application for Li-Ion Batteries. *Microporous Mesoporous Mater.* **2017**, *253*, 10–17. <https://doi.org/10.1016/j.micromeso.2017.06.033>.
- (150) Li, J.; Fu, L.; Xu, Z.; Zhu, J.; Yang, W.; Li, D.; Zhou, L. Electrochemical Properties of Carbon-Wrapped  $\text{FeF}_3$  Nanocomposite as Cathode Material for Lithium Ion Battery. *Electrochim. Acta* **2018**, *281*, 88–98. <https://doi.org/10.1016/J.ELECTACTA.2018.05.158>.
- (151) Tang, M.; Zhang, Z.; Wang, Z.; Liu, J.; Yan, H.; Peng, J. High-Temperature Electrochemical Performance of  $\text{FeF}_3/\text{C}$  Nanocomposite as a Cathode Material for Lithium-Ion Batteries. *J. Mater. Eng. Perform.* **2018**, *27*, 624–629. <https://doi.org/10.1007/s11665-018-3167-3>.
- (152) Fu, W.; Zhao, E.; Sun, Z.; Ren, X.; Magasinski, A.; Yushin, G. Iron Fluoride-Carbon Nanocomposite Nanofibers as Free-Standing Cathodes for High-Energy Lithium Batteries. *Adv. Funct. Mater.* **2018**, *28*, 1–8. <https://doi.org/10.1002/adfm.201801711>.
- (153) Guo, S. nan; Guo, H.; Wang, X.; Zhu, Y.; Hu, J.; Yang, M.; Zhao, L.; Wang, J. Iron Trifluoride as a High Voltage Cathode Material for Thermal Batteries. *J. Electrochem. Soc.* **2019**, *166*, A3599–A3605. <https://doi.org/10.1149/2.0371915jes>.
- (154) Takami, T.; Matsui, K.; Senoh, H.; Taguchi, N.; Shikano, M.; Sakaebe, H.; Fukunaga, T. Magnetic Behavior of Fe Nanoparticles Driven by Phase Transition of  $\text{FeF}_3$ . *J. Alloys Compd.* **2018**, *769*, 539–544. <https://doi.org/10.1016/j.jallcom.2018.08.040>.
- (155) Takami, T.; Matsui, K.; Senoh, H.; Shikano, M.; Sakaebe, H.; Fukunaga, T. Role of the Particle Size of Fe Nanoparticles in the Capacity of  $\text{FeF}_3$  Batteries. *AIP Adv.* **2019**, *9*, 45301. <https://doi.org/10.1063/1.5092144>.
- (156) Senoh, H.; Matsui, K.; Shikano, M.; Okumura, T.; Kiuchi, H.; Shimoda, K.; Yamanaka, K.; Ohta, T.; Fukunaga, T.; Sakaebe, H.; et al. Degradation Mechanism of Conversion-Type Iron Trifluoride: Toward Improvement of Cycle Performance. *ACS Appl. Mater. Interfaces* **2019**, *11*, 30959–30967. <https://doi.org/10.1021/acsami.9b10105>.
- (157) Wu, F.; Srot, V.; Chen, S.; Lorget, S.; Aken, P. A.; Maier, J.; Yu, Y. 3D Honeycomb Architecture Enables a High-Rate and Long-Life Iron (III) Fluoride–Lithium Battery. *Adv. Mater.* **2019**, *1905146*, 1905146. <https://doi.org/10.1002/adma.201905146>.

- (158) Eveillard, F.; Gervillié, C.; Taviot-Guého, C.; Leroux, F.; Guérin, K.; Sougrati, M. T.; Belin, S.; Delbègue, D. Unravelling Lithiation Mechanisms of Iron Trifluoride by Operando X-Ray Absorption Spectroscopy and MCR-ALS Chemometric Tools. *New J. Chem.* **2020**, *44*, 10153–10164. <https://doi.org/10.1039/c9nj06321h>.
- (159) Yang, Y.; Gao, L.; Shen, L.; Bao, N. Self-Assembled FeF<sub>3</sub> Nanocrystals Clusters Confined in Carbon Nanocages for High-Performance Li-Ion Battery Cathode. *J. Alloys Compd.* **2021**, 159799. <https://doi.org/10.1016/j.jallcom.2021.159799>.
- (160) Bai, Y.; Zhou, X.; Jia, Z.; Wu, C.; Yang, L.; Chen, M.; Zhao, H.; Wu, F.; Liu, G. Understanding the Combined Effects of Microcrystal Growth and Band Gap Reduction for Fe<sub>(1-x)</sub>Ti<sub>x</sub>F<sub>3</sub> Nanocomposites as Cathode Materials for Lithium-Ion Batteries. *Nano Energy* **2015**, *17*, 140–151. <https://doi.org/10.1016/j.nanoen.2015.08.006>.
- (161) Li, J.; Xu, S.; Huang, S.; Lu, L.; Lan, L.; Li, S. In Situ Synthesis of Fe<sub>1-x</sub>Co<sub>x</sub>F<sub>3</sub>/MWCNT Nanocomposites with Excellent Electrochemical Performance for Lithium-Ion Batteries. *J. Mater. Sci.* **2018**, *53*, 2697–2708. <https://doi.org/10.1007/s10853-017-1685-2>.
- (162) Su, J.; Nong, W.; Song, H.; Li, Y.; Wang, C. Enhanced Li-Storage Capability and Cyclability of Iron Fluoride Cathodes by Non-Equivalent Cobalt Doping. *J. Alloys Compd.* **2021**, 159395. <https://doi.org/https://doi.org/10.1016/j.jallcom.2021.159395>.
- (163) Groult, H.; Neveu, S.; Leclerc, S.; Porrás-Gutiérrez, A. G.; Julien, C. M.; Tressaud, A.; Durand, E.; Penin, N.; Labrugere, C. Nano-CoF<sub>3</sub> Prepared by Direct Fluorination with F<sub>2</sub> Gas: Application as Electrode Material in Li-Ion Battery. *J. Fluor. Chem.* **2017**, *196*, 117–127. <https://doi.org/10.1016/j.jfluchem.2016.10.003>.
- (164) Kitajou, A.; Eguchi, K.; Ishido, Y.; Setoyama, H.; Okajima, T.; Okada, S. Electrochemical Properties of Titanium Fluoride with High Rate Capability for Lithium-Ion Batteries. *J. Power Sources* **2019**, *419*, 1–5. <https://doi.org/10.1016/j.jpowsour.2019.02.056>.
- (165) Li, H.; Richter, G.; Maier, J. Reversible Formation and Decomposition of LiF Clusters Using Transition Metal Fluorides as Precursors and Their Application in Rechargeable Li Batteries. *Adv. Mater.* **2003**, *15*, 736–739. <https://doi.org/10.1002/adma.200304574>.
- (166) Owen, N.; Zhang, Q. Investigations of Aluminum Fluoride as a New Cathode Material for Lithium-Ion Batteries. *J. Appl. Electrochem.* **2017**, *47*, 417–431. <https://doi.org/10.1007/s10800-017-1049-2>.
- (167) Pérez-Flores, J. C.; Villamor, R.; Ávila-Brandé, D.; Gallardo Amores, J. M.; Morán, E.; Kuhn, A.; García-Alvarado, F. VO<sub>2</sub>F: A New Transition Metal Oxyfluoride with High Specific Capacity for Li Ion Batteries. *J. Mater. Chem. A* **2015**, *3*, 20508–20515. <https://doi.org/10.1039/c5ta05434f>.
- (168) Wang, X.; Lin, Y. C.; Zhou, H.; Omenya, F.; Chu, I. H.; Karki, K.; Sallis, S.; Rana, J.; Piper, L. F. J.; Chernova, N. A.; et al. Structural Changes in a High-Energy Density VO<sub>2</sub>F Cathode upon Heating and Li Cycling. *ACS Appl. Energy Mater.* **2018**, *1*, 4514–4521. <https://doi.org/10.1021/acsaem.8b00473>.

- (169) Chen, R.; Maawad, E.; Knapp, M.; Ren, S.; Beran, P.; Witter, R.; Hempelmann, R. Lithiation-Driven Structural Transition of VO<sub>2</sub>F into Disordered Rock-Salt Li<sub>x</sub>VO<sub>2</sub>F. *RSC Adv.* **2016**, *6*, 65112–65118. <https://doi.org/10.1039/C6RA14276A>.
- (170) Cambaz, M. A.; Vinayan, B. P.; Clemens, O.; Munnangi, A. R.; Chakravadhanula, V. S. K.; Kübel, C.; Fichtner, M. Vanadium Oxyfluoride/Few-Layer Graphene Composite as a High-Performance Cathode Material for Lithium Batteries. *Inorg. Chem.* **2016**, *55*, 3789–3796. <https://doi.org/10.1021/acs.inorgchem.5b02687>.
- (171) Wang, X.; Huang, Y.; Ji, D.; Omenya, F.; Karki, K.; Sallis, S.; Piper, L. F. J.; Wiaderek, K. M.; Chapman, K. W.; Chernova, N. A.; et al. Structure Evolution and Thermal Stability of High-Energy-Density Li-Ion Battery Cathode Li<sub>2</sub>VO<sub>2</sub>F. *J. Electrochem. Soc.* **2017**, *164*, A1552–A1558. <https://doi.org/10.1149/2.1071707jes>.
- (172) Ren, S.; Chen, R.; Maawad, E.; Dolotko, O.; Guda, A. A.; Shapovalov, V.; Wang, D.; Hahn, H.; Fichtner, M. Improved Voltage and Cycling for Li<sup>+</sup> Intercalation in High-Capacity Disordered Oxyfluoride Cathodes. *Adv. Sci.* **2015**, *2*, 1500128. <https://doi.org/10.1002/advs.201500128>.
- (173) Chen, R.; Ren, S.; Mu, X.; Maawad, E.; Zander, S.; Hempelmann, R.; Hahn, H. High-Performance Low-Temperature Li<sup>+</sup> Intercalation in Disordered Rock-Salt Li-Cr-V Oxyfluorides. *ChemElectroChem* **2016**, *3*, 892–895. <https://doi.org/10.1002/celec.201600033>.
- (174) Wu, W.; Wang, X.; Wang, X.; Wang, G. B.; Yang, S. Y.; Wei, J. L.; Li, N. Studies on Preparation and Electrochemical Performances of FeF<sub>3</sub>(H<sub>2</sub>O)<sub>0.33</sub> Cathode Material for the Application of Lithium Rechargeable Battery. *J. Funct. Mater.* **2008**, *39*, 1824–1827.
- (175) Li, C.; Gu, L.; Tsukimoto, S.; van Aken, P. A.; Maier, J. Low-Temperature Ionic-Liquid-Based Synthesis of Nanostructured Iron-Based Fluoride Cathodes for Lithium Batteries. *Adv. Mater.* **2010**, *22*, 3650–3654. <https://doi.org/10.1002/adma.201000535>.
- (176) Li, C.; Gu, L.; Tong, J.; Tsukimoto, S.; Maier, J. A Mesoporous Iron-Based Fluoride Cathode of Tunnel Structure for Rechargeable Lithium Batteries. *Adv. Funct. Mater.* **2011**, *21*, 1391–1397. <https://doi.org/10.1002/adfm.201002213>.
- (177) Li, C.; Gu, L.; Tong, J.; Maier, J. Carbon Nanotube Wiring of Electrodes for High-Rate Lithium Batteries Using an Imidazolium-Based Ionic Liquid Precursor as Dispersant and Binder: A Case Study on Iron Fluoride Nanoparticles. *ACS Nano* **2011**, *5*, 2930–2938. <https://doi.org/10.1021/nn1035608>.
- (178) Liu, L.; Zhou, M.; Yi, L.; Guo, H.; Tan, J.; Shu, H.; Yang, X.; Yang, Z.; Wang, X. Excellent Cycle Performance of Co-Doped FeF<sub>3</sub>/C Nanocomposite Cathode Material for Lithium-Ion Batteries. *J. Mater. Chem.* **2012**, *22*, 17539–17550. <https://doi.org/10.1039/C2JM32936K>.
- (179) Liu, L.; Zhou, M.; Wang, X.; Yang, Z.; Tian, F.; Wang, X. Synthesis and Electrochemical Performance of Spherical FeF<sub>3</sub>/ACMB Composite as Cathode Material for Lithium-Ion Batteries. *J. Mater. Sci.* **2012**, *47*, 1819–1824. <https://doi.org/10.1007/s10853-011-5968-8>.
- (180) Li, C.; Yin, C.; Mu, X.; Maier, J. Top-Down Synthesis of Open Framework Fluoride for Lithium



- and Sodium Batteries. *Chem. Mater.* **2013**, *25*, 962–969. <https://doi.org/10.1021/cm304127c>.
- (181) Li, B.; Rooney, D. W.; Zhang, N.; Sun, K. An in Situ Ionic-Liquid-Assisted Synthetic Approach to Iron Fluoride/Graphene Hybrid Nanostructures as Superior Cathode Materials for Lithium Ion Batteries. *ACS Appl. Mater. Interfaces* **2013**, *5*, 5057–5063. <https://doi.org/10.1021/am400873e>.
- (182) Xu, X.; Chen, S.; Shui, M.; Xu, L.; Zheng, W.; Shu, J.; Cheng, L.; Feng, L.; Ren, Y. One Step Solid State Synthesis of  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}/\text{C}$  Nano-Composite as Cathode Material for Lithium-Ion Batteries. *Ceram. Int.* **2014**, *40*, 3145–3148. <https://doi.org/10.1016/J.CERAMINT.2013.09.130>.
- (183) Liu, J.; Liu, W.; Ji, S.; Wan, Y.; Gu, M.; Yin, H.; Zhou, Y. Iron Fluoride Hollow Porous Microspheres: Facile Solution-Phase Synthesis and Their Application for Li-Ion Battery Cathodes. *Chem. - A Eur. J.* **2014**, *20*, 5815–5820. <https://doi.org/10.1002/chem.201304713>.
- (184) Di Carlo, L.; Conte, D. E.; Kemnitz, E.; Pinna, N. Microwave-Assisted Fluorolytic Sol–Gel Route to Iron Fluoride Nanoparticles for Li-Ion Batteries. *Chem. Commun.* **2014**, *50*, 460–462. <https://doi.org/10.1039/C3CC47413E>.
- (185) Li, B.; Zhang, N.; Sun, K. Confined Iron Fluoride@CMK-3 Nanocomposite as an Ultrahigh Rate Capability Cathode for Li-Ion Batteries. *Small* **2014**, *10*, 2039–2046. <https://doi.org/10.1002/smll.201303375>.
- (186) Li, B.; Cheng, Z.; Zhang, N.; Sun, K. Self-Supported, Binder-Free 3D Hierarchical Iron Fluoride Flower-like Array as High Power Cathode Material for Lithium Batteries. *Nano Energy* **2014**, *4*, 7–13. <https://doi.org/10.1016/j.nanoen.2013.12.003>.
- (187) Tan, J.; Liu, L.; Hu, H.; Yang, Z.; Guo, H.; Wei, Q.; Yi, X.; Yan, Z.; Zhou, Q.; Huang, Z.; et al. Iron Fluoride with Excellent Cycle Performance Synthesized by Solvothermal Method as Cathodes for Lithium Ion Batteries. *J. Power Sources* **2014**, *251*, 75–84. <https://doi.org/10.1016/j.jpowsour.2013.11.004>.
- (188) Long, Z.; Hu, W.; Liu, L.; Qiu, G.; Qiao, W.; Guan, X.; Qiu, X. Mesoporous Iron Trifluoride Microspheres as Cathode Materials for Li-Ion Batteries. *Electrochim. Acta* **2015**, *151*, 355–362. <https://doi.org/10.1016/j.electacta.2014.11.029>.
- (189) Fan, L.; Li, B.; Zhang, N.; Sun, K. Carbon Nanohorns Carried Iron Fluoride Nanocomposite with Ultrahigh Rate Lithium Ion Storage Properties. *Sci. Rep.* **2015**, *5*, 1–9. <https://doi.org/10.1038/srep12154>.
- (190) Hu, X.; Ma, M.; Mendes, R. G.; Zeng, M.; Zhang, Q.; Xue, Y.; Zhang, T.; Rummeli, M. H.; Fu, L. Li-Storage Performance of Binder-Free and Flexible Iron Fluoride@graphene Cathodes. *J. Mater. Chem. A* **2015**, *3*, 23930–23935. <https://doi.org/10.1039/c5ta08014b>.
- (191) Wei, S.; Wang, X.; Jiang, M.; Zhang, R.; Shen, Y.; Hu, H. The  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}/\text{C}$  Nanocomposite with Open Mesoporous Structure as High-Capacity Cathode Material for Lithium/Sodium Ion Batteries. *J. Alloys Compd.* **2016**, *689*, 945–951. <https://doi.org/10.1016/j.jallcom.2016.08.080>.

- (192) Rao, R. S.; Pralong, V.; Varadaraju, U. V. Facile Synthesis and Reversible Lithium Insertion Studies on Hydrated Iron Trifluoride  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$ . *Solid State Sci.* **2016**, *55*, 77–82. <https://doi.org/10.1016/j.solidstatesciences.2016.02.008>.
- (193) Pohl, A.; Faraz, M.; Schröder, A.; Baunach, M.; Schabel, W.; Guda, A.; Shapovalov, V.; Soldatov, A.; Chakravadhanula, V. S. K.; Kübel, C.; et al. Development of a Water Based Process for Stable Conversion Cathodes on the Basis of  $\text{FeF}_3$ . *J. Power Sources* **2016**, *313*, 213–222. <https://doi.org/10.1016/j.jpowsour.2016.02.080>.
- (194) Han, Y.; Li, H.; Li, J.; Si, H.; Zhu, W.; Qiu, X. Hierarchical Mesoporous Iron Fluoride with Superior Rate Performance for Lithium-Ion Batteries. *ACS Appl. Mater. Interfaces* **2016**, *8*, 32869–32874. <https://doi.org/10.1021/acsami.6b11889>.
- (195) Yang, J.; Xu, Z.; Sun, H.; Zhou, X. A Three-Dimensional Interlayer Composed of Graphene and Porous Carbon for Long-Life, High Capacity Lithium-Iron Fluoride Battery. *Electrochim. Acta* **2016**, *220*, 75–82. <https://doi.org/10.1016/j.electacta.2016.10.076>.
- (196) Hu, J.; Zhang, Y.; Cao, D.; Li, C. Dehydrating Bronze Iron Fluoride as a High Capacity Conversion Cathode for Lithium Batteries. *J. Mater. Chem. A* **2016**, *4*, 16166–16174. <https://doi.org/10.1039/c6ta05929e>.
- (197) Bai, Y.; Zhou, X.; Zhan, C.; Ma, L.; Yuan, Y.; Wu, C.; Chen, M.; Chen, G.; Ni, Q.; Wu, F.; et al. 3D Hierarchical Nano-Flake/Micro-Flower Iron Fluoride with Hydration Water Induced Tunnels for Secondary Lithium Battery Cathodes. *Nano Energy* **2017**, *32*, 10–18. <https://doi.org/10.1016/j.nanoen.2016.12.017>.
- (198) Li, L.; Zhu, J.; Xu, M.; Jiang, J.; Li, C. M. In Situ Engineering Toward Core Regions: A Smart Way to Make Applicable  $\text{FeF}_3$ @Carbon Nanoreactor Cathodes for Li-Ion Batteries. *ACS Appl. Mater. Interfaces* **2017**, *9*, 17992–18000. <https://doi.org/10.1021/acsami.7b04256>.
- (199) Li, Y.; Zhou, X.; Bai, Y.; Chen, G.; Wang, Z.; Li, H.; Wu, F.; Wu, C. Building an Electronic Bridge via Ag Decoration to Enhance Kinetics of Iron Fluoride Cathode in Lithium-Ion Batteries. *ACS Appl. Mater. Interfaces* **2017**, *9*, 19852–19860. <https://doi.org/10.1021/acsami.7b03980>.
- (200) Qiu, D.; Fu, L.; Zhan, C.; Lu, J.; Wu, D. Seeding Iron Trifluoride Nanoparticles on Reduced Graphite Oxide for Lithium-Ion Batteries with Enhanced Loading and Stability. *ACS Appl. Mater. Interfaces* **2018**, *10*, 29505–29510. <https://doi.org/10.1021/acsami.8b08526>.
- (201) Zhai, J.; Lei, Z.; Rooney, D.; Wang, H.; Sun, K. Self-Templated Fabrication of Micro/Nano Structured Iron Fluoride for High-Performance Lithium-Ion Batteries. *J. Power Sources* **2018**, *396*, 371–378. <https://doi.org/10.1016/j.jpowsour.2018.06.048>.
- (202) Zhang, R.; Wang, X. X.; Wang, X. X.; Liu, M.; Wei, S.; Wang, Y.; Hu, H. Iron Fluoride Packaged into 3D Order Mesoporous Carbons as High-Performance Sodium-Ion Battery Cathode Material. *J. Electrochem. Soc.* **2018**, *165*, A89–A96. <https://doi.org/10.1149/2.0421802jes>.
- (203) Liu, M.; Liu, L.; Hu, H.; Yang, L.; Yang, Z.; Wang, Y.; Wang, X. Flowerlike Mesoporous

FeF<sub>3</sub>·0.33H<sub>2</sub>O with 3D Hierarchical Nanostructure: Size-Controlled Green-Synthesis and Application as Cathodes for Na-Ion Batteries. *ACS Appl. Energy Mater.* **2018**, *1*, 7153–7163. <https://doi.org/10.1021/acsaem.8b01585>.

- (204) Tang, Y.; An, J.; Xing, H.; Wang, X.; Zhai, B.; Zhang, F.; Song, Y.; Li, G. Synthesis of Iron-Fluoride Materials with Controlled Nanostructures and Composition through a Template-Free Solvothermal Route for Lithium Ion Batteries. *New J. Chem.* **2018**, *42*, 9091–9097. <https://doi.org/10.1039/c8nj00932e>.
- (205) Liu, M.; Wang, X.; Zhang, R.; Liu, L.; Hu, H.; Wang, Y.; Wei, S. Hollow Porous FeF<sub>3</sub>·0.33H<sub>2</sub>O Microspheres by AlPO<sub>4</sub> Coating as a Cathode Material of Na-Ion Batteries. *J. Energy Storage* **2018**, *18*, 103–111. <https://doi.org/10.1016/j.est.2018.04.026>.
- (206) Wei, S.; Wang, X.; Liu, M.; Zhang, R.; Wang, G.; Hu, H. Spherical FeF<sub>3</sub>·0.33H<sub>2</sub>O /MWCNTs Nanocomposite with Mesoporous Structure as Cathode Material of Sodium Ion Battery. *J. Energy Chem.* **2018**, *27*, 573–581. <https://doi.org/10.1016/j.jechem.2017.10.032>.
- (207) Bensalah, N.; Mustafa, N. In Situ Generated MWCNT-FeF<sub>3</sub>·0.33H<sub>2</sub>O Nanocomposites toward Stable Performance Cathode Material for Lithium Ion Batteries. *Emergent Mater.* **2019**, *2*, 59–66. <https://doi.org/10.1007/s42247-018-00021-5>.
- (208) Zhou, H.; Sun, H.; Wang, T.; Gao, Y.; Ding, J.; Xu, Z.; Tang, J.; Jia, M.; Yang, J.; Zhu, J. Low Temperature Nanotailoring of Hydrated Compound by Alcohols: FeF<sub>3</sub>·3H<sub>2</sub>O as an Example. Preparation of Nanosized FeF<sub>3</sub>·0.33H<sub>2</sub>O Cathode Material for Li-Ion Batteries. *Inorg. Chem.* **2019**, *58*, 6765–6771. <https://doi.org/10.1021/acs.inorgchem.9b00054>.
- (209) Zhang, Q.; Liu, N. N.; Sun, C. Z.; Fan, L. S.; Zhang, N. Q.; Sun, K. N. Ultrasmall Iron Fluoride Nanoparticles Embedded in Graphitized Porous Carbon Derived from Fe-Based Metal Organic Frameworks as High-Performance Cathode Materials for Li Batteries. *ChemElectroChem* **2019**, *6*, 2189–2194. <https://doi.org/10.1002/celec.201900244>.
- (210) Zhang, Q.; Wu, X.; Gong, S.; Fan, L.; Zhang, N. Iron Fluoride Nanoparticles Embedded in a Nitrogen and Oxygen Dual-Doped 3D Porous Carbon Derived from Nori for High Rate Cathode in Lithium-Ion Battery. *ChemistrySelect* **2019**, *4*, 10334–10339. <https://doi.org/10.1002/slct.201902478>.
- (211) Chen, G.; Zhou, X.; Bai, Y.; Yuan, Y.; Li, Y.; Chen, M.; Ma, L.; Tan, G.; Hu, J.; Wang, Z.; et al. Enhanced Lithium Storage Capability of FeF<sub>3</sub>·0.33H<sub>2</sub>O Single Crystal with Active Insertion Site Exposed. *Nano Energy* **2019**, *56*, 884–892. <https://doi.org/10.1016/j.nanoen.2018.11.080>.
- (212) Zhai, J.; Lei, Z.; Rooney, D.; Sun, K. Top-down Synthesis of Iron Fluoride/Reduced Graphene Nanocomposite for High Performance Lithium-Ion Battery. *Electrochim. Acta* **2019**, *313*, 497–504. <https://doi.org/10.1016/j.electacta.2019.04.024>.
- (213) Lu, L.; Li, S.; Li, J.; Lan, L.; Lu, Y.; Xu, S.; Huang, S.; Pan, C.; Zhao, F. High-Performance Cathode Material of FeF<sub>3</sub>·0.33H<sub>2</sub>O Modified with Carbon Nanotubes and Graphene for Lithium-Ion Batteries. *Nanoscale Res. Lett.* **2019**, *14*, 100. <https://doi.org/10.1186/s11671->

019-2925-y.

- (214) Lin, J.; Zhu, L.; Chen, S.; Li, Q.; He, Z.; Cai, Z.; Cao, L.; Yuan, Z.; Liu, J. Self-Templated Formation of Hollow Yolk-Like Spheres Iron Fluoride as Cathode Material for High-Performance Li-Ion Batteries. *J. Electrochem. Soc.* **2019**, *166*, A2074–A2082. <https://doi.org/10.1149/2.0991910jes>.
- (215) Zhang, Q.; Sun, C.; Fan, L.; Zhang, N.; Sun, K. Iron Fluoride Vertical Nanosheets Array Modified with Graphene Quantum Dots as Long-Life Cathode for Lithium Ion Batteries. *Chem. Eng. J.* **2019**, *371*, 245–251. <https://doi.org/10.1016/j.cej.2019.04.073>.
- (216) Zhang, C.; An, S.; Li, W.; Xu, H.; Hao, W.; Liu, W.; Li, Z.; Qiu, X. Hierarchical Mesoporous Iron Fluoride and Reduced Graphene Oxide Nanocomposite as Cathode Materials for High-Performance Sodium-Ion Batteries. *ACS Appl. Mater. Interfaces* **2020**, *12*, 17538–17546. <https://doi.org/10.1021/acsami.0c01652>.
- (217) Zhang, Q.; Zhang, Y.; Yin, Y.; Fan, L.; Zhang, N. Packing  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  into Porous Graphene/Carbon Nanotube Network as High Volumetric Performance Cathode for Lithium Ion Battery. *J. Power Sources* **2020**, *447*, 227303. <https://doi.org/10.1016/j.jpowsour.2019.227303>.
- (218) Chen, S.; Lin, J.; Shi, Q.; Cai, Z.; Cao, L.; Zhu, L.; Yuan, Z. Nanoscale Iron Fluoride Supported by Three-Dimensional Porous Graphene as Long-Life Cathodes for Lithium-Ion Batteries. *J. Electrochem. Soc.* **2020**, *167*, 080506. <https://doi.org/10.1149/1945-7111/ab88be>.
- (219) Ding, J.; Zhou, X.; Wang, Q.; Luo, C.; Yang, J.; Tang, J. N, S Co-Doped Porous Carbon from Antibiotic Bacteria Residues Enables a High-Performance  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  Cathode for Li-Ion Batteries. *ChemElectroChem* **2020**, *7*, 4931–4935. <https://doi.org/10.1002/celec.202001353>.
- (220) Lin, J.; Chen, S.; Zhu, L.; Yuan, Z.; Liu, J. Soft-Template Fabrication of Hierarchical Nanoparticle Iron Fluoride as High-Capacity Cathode Materials for Li-Ion Batteries. *Electrochim. Acta* **2020**, *364*, 137293. <https://doi.org/10.1016/j.electacta.2020.137293>.
- (221) Cheng, Q.; Pan, Y.; Chen, Y.; Zeb, A.; Lin, X.; Yuan, Z.; Liu, J. Nanostructured Iron Fluoride Derived from Fe-Based Metal-Organic Framework for Lithium Ion Battery Cathodes. *Inorg. Chem.* **2020**, *59*, 12700–12710. <https://doi.org/10.1021/acs.inorgchem.0c01783>.
- (222) Liu, M.; Liu, L.; Li, M.; Chen, B.; Lei, H.; Hu, H.; Wang, X. Preparation and Li/Na Ion Storage Performance of Raspberry-like Hierarchical  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  Micro-Sized Spheres with Controllable Morphology. *J. Alloys Compd.* **2020**, *829*, 154215. <https://doi.org/10.1016/j.jallcom.2020.154215>.
- (223) Zhang, L.; Litao, Y.; Li, O. L.; Choi, S.-Y.; Saeed, G.; Kim, K. H.  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$ @Carbon Nanosheets with Honeycomb Architectures for High-Capacity Lithium-Ion Cathode Storage by Enhanced Pseudocapacitance. *J. Mater. Chem. A* **2021**. <https://doi.org/10.1039/d1ta03141d>.
- (224) Zeng, C.; Chen, F.; Ye, Q.; Guo, Q.; Li, C.; Huang, C. Facile Preparation of Hierarchical Micronano  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  by a One-Pot Method with Dual Surfactants. *Nanotechnology* **2021**,

32, 155402. <https://doi.org/10.1088/1361-6528/abd6d0>.

- (225) Wu, W.; Wang, Y.; Wang, X.; Chen, Q.; Wang, X.; Yang, S.; Liu, X.; Guo, J.; Yang, Z. Structure and Electrochemical Performance of  $\text{FeF}_3/\text{V}_2\text{O}_5$  Composite Cathode Material for Lithium-Ion Battery. *J. Alloys Compd.* **2009**, *486*, 93–96. <https://doi.org/10.1016/j.jallcom.2009.07.063>.
- (226) Wu, W.; Wang, X.; Wang, X.; Yang, S.; Liu, X.; Chen, Q. Effects of  $\text{MoS}_2$  Doping on the Electrochemical Performance of  $\text{FeF}_3$  Cathode Materials for Lithium-Ion Batteries. *Mater. Lett.* **2009**, *63*, 1788–1790. <https://doi.org/10.1016/j.matlet.2009.05.041>.
- (227) Zhou, X.; Ding, J.; Tang, J.; Yang, J.; Wang, H.; Jia, M. Tailored  $\text{MoO}_3$ -Encapsulated  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  Composites as High Performance Cathodes for Li-Ion Batteries. *J. Electroanal. Chem.* **2019**, *847*, 113227. <https://doi.org/10.1016/j.jelechem.2019.113227>.
- (228) Yang, J.; Xu, Z.; Zhou, H.; Tang, J.; Sun, H.; Ding, J.; Zhou, X. A Cathode Material Based on the Iron Fluoride with an Ultra-Thin  $\text{Li}_3\text{FeF}_6$  Protective Layer for High-Capacity Li-Ion Batteries. *J. Power Sources* **2017**, *363*, 244–250. <https://doi.org/10.1016/J.JPOWSOUR.2017.07.091>.
- (229) Zhang, R.; Wang, X.; Wei, S.; Wang, X.; Liu, M.; Hu, H. Iron Fluoride Microspheres by Titanium Dioxide Surface Modification as High Capacity Cathode of Li-Ion Batteries. *J. Alloys Compd.* **2017**, *719*, 331–340. <https://doi.org/10.1016/j.jallcom.2017.05.185>.
- (230) Duttine, M.; Dambournet, D.; Penin, N.; Carlier, D.; Bourgeois, L.; Wattiaux, A.; Chapman, K. W.; Chupas, P. J.; Groult, H.; Durand, E.; et al. Tailoring the Composition of a Mixed Anion Iron-Based Fluoride Compound: Evidence for Anionic Vacancy and Electrochemical Performance in Lithium Cells. *Chem. Mater.* **2014**, *26*, 4190–4199. <https://doi.org/10.1021/cm501396n>.
- (231) Dambournet, D.; Chapman, K. W.; Duttine, M.; Borkiewicz, O.; Chupas, P. J.; Groult, H. Lithium Insertion Mechanism in Iron-Based Oxyfluorides with Anionic Vacancies Probed by PDF Analysis. *ChemistryOpen* **2015**, *4*, 443–447. <https://doi.org/10.1002/open.201500031>.
- (232) Burbano, M.; Duttine, M.; Morgan, B. J.; Borkiewicz, O. J.; Chapman, K. W.; Wattiaux, A.; Demourgues, A.; Groult, H.; Salanne, M.; Dambournet, D. Impact of Anion Vacancies on the Local and Electronic Structures of Iron-Based Oxyfluoride Electrodes. *J. Phys. Chem. Lett.* **2019**, *10*, 107–112. <https://doi.org/10.1021/acs.jpcclett.8b03503>.
- (233) Lemoine, K.; Zhang, L.; Dambournet, D.; Greneche, J. M.; Hemon-Ribaud, A.; Leblanc, M.; Borkiewicz, O. J.; Tarascon, J.-M.; Maisonneuve, V.; Lhoste, J. Synthesis by Thermal Decomposition of Two Iron Hydroxy Fluorides : Structural Effects of Li Insertion. *Chem. Mater.* **2019**, *31*, 4246–4257. <https://doi.org/10.1021/acs.chemmater.9b01252>.
- (234) Wei, S.; Wang, X.; Yu, R.; Zhang, R.; Liu, M.; Yang, Z.; Hu, H. Ti-Doped  $\text{Fe}_{1-x}\text{Ti}_x\text{F}_3 \cdot 0.33\text{H}_2\text{O}/\text{C}$  Nanocomposite as an Ultrahigh Rate Capability Cathode Materials of Lithium Ion Batteries. *J. Alloys Compd.* **2017**, *702*, 372–380. <https://doi.org/10.1016/j.jallcom.2017.01.240>.
- (235) Yang, Z.; Zhang, Z.; Yuan, Y.; Huang, Y.; Wang, X.; Chen, X.; Wei, S. First-Principles Study of Ti Doping in  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$ . *Curr. Appl. Phys.* **2016**, *16*, 905–913. <https://doi.org/10.1016/j.cap.2016.05.010>.

- (236) Ding, J.; Zhou, X.; Wang, H.; Yang, J.; Gao, Y.; Tang, J. Mn-Doped  $\text{Fe}_{1-x}\text{Mn}_x\text{F}_3 \cdot 0.33\text{H}_2\text{O}/\text{C}$  Cathodes for Li-Ion Batteries: First-Principles Calculations and Experimental Study. *ACS Appl. Mater. Interfaces* **2019**, *11*, 3852–3860. <https://doi.org/10.1021/acsami.8b17069>.
- (237) Lu, Y.; Huang, S.; Zhang, Z.; Huang, X.; Lan, L.; Lu, L.; Li, S.; Li, J.; Pan, C.; Zhao, F. Mn-Doped  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  with Enhanced Electrochemical Performance as Cathode Materials for Lithium-Ion Batteries. *Ionics (Kiel)*. **2019**, 1–8. <https://doi.org/10.1007/s11581-019-03094-2>.
- (238) Liu, M.; Wang, Q.; Chen, B.; Lei, H.; Liu, L.; Wu, C.; Wang, X.; Yang, Z. Band-Gap  $\text{FeF}_3 \cdot 0.33\text{H}_2\text{O}$  Nanosphere via Ni Doping as a High-Performance Lithium-Ion Battery Cathode. *ACS Sustain. Chem. Eng.* **2020**, *8*, 15651–15660. <https://doi.org/10.1021/acssuschemeng.0c05258>.
- (239) Lemoine, K.; Moury, R.; Lhoste, J.; Ribaud, A.; Leblanc, M.; Grenèche, J.-M.; Tarascon, J.-M.; Maisonneuve, V. Stabilization of a Mixed Iron Vanadium Based Hexagonal Tungsten Bronze Hydroxyfluoride  $\text{HTB}-(\text{Fe}_{0.55}\text{V}_{0.45})\text{F}_{2.67}(\text{OH})_{0.33}$  as a Positive Electrode for Lithium-Ion Batteries. *Dalt. Trans.* **2020**, *49*, 8186–8193. <https://doi.org/10.1039/d0dt01310b>.
- (240) Li, C.; Yin, C.; Gu, L.; Dinnebier, R. E.; Mu, X.; Van Aken, P. A.; Maier, J. An  $\text{FeF}_3 \cdot 0.5\text{H}_2\text{O}$  Polytype: A Microporous Framework Compound with Intersecting Tunnels for Li and Na Batteries. *J. Am. Chem. Soc.* **2013**, *135*, 11425–11428. <https://doi.org/10.1021/ja402061q>.
- (241) Li, C.; Mu, X.; Van Aken, P. A.; Maier, J. A High-Capacity Cathode for Lithium Batteries Consisting of Porous Microspheres of Highly Amorphized Iron Fluoride Densified from Its Open Parent Phase. *Adv. Energy Mater.* **2013**, *3*, 113–119. <https://doi.org/10.1002/aenm.201200209>.
- (242) Lu, Y.; Wen, Z.; Rui, K.; Wu, X.; Cui, Y. Worm-like Mesoporous Structured Iron-Based Fluoride: Facile Preparation and Application as Cathodes for Rechargeable Lithium Ion Batteries. *J. Power Sources* **2013**, *244*, 306–311. <https://doi.org/10.1016/j.jpowsour.2013.01.035>.
- (243) Chu, Q.; Xing, Z.; Tian, J.; Ren, X.; Asiri, A. M.; Al-Youbi, A. O.; Alamry, K. A.; Sun, X. Facile Preparation of Porous  $\text{FeF}_3$  Nanospheres as Cathode Materials for Rechargeable Lithium-Ion Batteries. *J. Power Sources* **2013**, *236*, 188–191. <https://doi.org/10.1016/j.jpowsour.2013.02.026>.
- (244) Lu, Y.; Wen, Z.; Jin, J.; Rui, K.; Wu, X. Hierarchical Mesoporous Iron-Based Fluoride with Partially Hollow Structure: Facile Preparation and High Performance as Cathode Material for Rechargeable Lithium Ion Batteries. *Phys. Chem. Chem. Phys.* **2014**, *16*, 8556–8562. <https://doi.org/10.1039/c4cp00568f>.
- (245) Jiang, M.; Wang, X.; Shen, Y.; Hu, H.; Fu, Y.; Yang, X. New Iron-Based Fluoride Cathode Material Synthesized by Non-Aqueous Ionic Liquid for Rechargeable Sodium Ion Batteries. *Electrochim. Acta* **2015**, *186*, 7–15. <https://doi.org/10.1016/j.electacta.2015.10.159>.
- (246) Wei, S.; Wang, X.; Zhang, R.; Hu, H.; Shen, Y.; Liu, J. Preparation and Performance of Spherical  $\text{FeF}_{2.5} \cdot 0.5\text{H}_2\text{O}$  Nanoparticles Wrapped by MWCNTs as Cathode Material of Lithium

- Ion Batteries. *RSC Adv.* **2016**, *6*, 97759–97769. <https://doi.org/10.1039/c6ra20314k>.
- (247) Jiang, M.; Wang, X.; Hu, H.; Wei, S.; Fu, Y.; Shen, Y. In Situ Growth and Performance of Spherical  $\text{Fe}_2\text{F}_5 \cdot \text{H}_2\text{O}$  Nanoparticles in Multi-Walled Carbon Nanotube Network Matrix as Cathode Material for Sodium Ion Batteries. *J. Power Sources* **2016**, *316*, 170–175. <https://doi.org/10.1016/j.jpowsour.2016.03.007>.
- (248) Jiang, M.; Wang, X.; Wei, S.; Shen, Y.; Hu, H. An Ionic-Liquid-Assisted Approach to Synthesize a Reduced Graphene Oxide Loading Iron-Based Fluoride as a Cathode Material for Sodium-Ion Batteries. *J. Alloys Compd.* **2016**, *670*, 362–368. <https://doi.org/10.1016/j.jallcom.2016.01.248>.
- (249) Ali, G.; Lee, J.-H.; Cho, B. W.; Nam, K.-W.; Ahn, D.; Chang, W.; Oh, S. H.; Chung, K. Y. Probing the Sodiation-Desodiation Reactions in Nano-Sized Iron Fluoride Cathode. *Electrochim. Acta* **2016**, *191*, 307–316. <https://doi.org/10.1016/j.electacta.2016.01.080>.
- (250) Rao, R. S.; Pralong, V.; Varadaraju, U. V. Facile Synthesis and Lithium Reversible Insertion on Iron Hydrated Trifluorides  $\text{FeF}_3 \cdot 0.5\text{H}_2\text{O}$ . *Mater. Lett.* **2016**, *170*, 130–134. <https://doi.org/10.1016/j.matlet.2016.02.008>.
- (251) Li, X.; Zhang, Y.; Meng, Y.; Tan, G.; Ou, J.; Wang, Y.; Zhao, Q.; Yuan, H.; Xiao, D. Three-Dimensional Nanocomposite of Iron-Based Fluoride Loaded in N-Doped Porous Carbon as a High-Performance Cathode for Rechargeable Li-Ion Batteries. *ChemElectroChem* **2017**, *4*, 1856–1862. <https://doi.org/10.1002/celec.201700259>.
- (252) Ali, G.; Lee, J. H.; Chang, W.; Cho, B. W.; Jung, H. G.; Nam, K. W.; Chung, K. Y. Lithium Intercalation Mechanism into  $\text{FeF}_3 \cdot 0.5\text{H}_2\text{O}$  as a Highly Stable Composite Cathode Material. *Sci. Rep.* **2017**, *7*, 42237. <https://doi.org/10.1038/srep42237>.
- (253) Ali, G.; Rahman, G.; Chung, K. Y. Cobalt-Doped Pyrochlore-Structured Iron Fluoride as a Highly Stable Cathode Material for Lithium-Ion Batteries. *Electrochim. Acta* **2017**, *238*, 49–55. <https://doi.org/10.1016/j.electacta.2017.04.006>.
- (254) Liu, M.; Wang, X.; Wei, S.; Hu, H.; Zhang, R.; Liu, L. Cr-Doped  $\text{Fe}_2\text{F}_5 \cdot \text{H}_2\text{O}$  with Open Framework Structure as a High Performance Cathode Material of Sodium-Ion Batteries. *Electrochim. Acta* **2018**, *269*, 479–489. <https://doi.org/10.1016/j.electacta.2018.02.159>.
- (255) Han, Y.; Hu, J.; Yin, C.; Zhang, Y.; Xie, J.; Yin, D.; Li, C. Iron-Based Fluorides of Tetragonal Tungsten Bronze Structure as Potential Cathodes for Na-Ion Batteries. *J. Mater. Chem. A* **2016**, *4*, 7382–7389. <https://doi.org/10.1039/C6TA02061E>.
- (256) Ma, D. L.; Cao, Z. Y.; Wang, H. G.; Huang, X. L.; Wang, L. M.; Zhang, X. B. Three-Dimensionally Ordered Macroporous  $\text{FeF}_3$  and Its in Situ Homogenous Polymerization Coating for High Energy and Power Density Lithium Ion Batteries. *Energy Environ. Sci.* **2012**, *5*, 8538–8542. <https://doi.org/10.1039/c2ee22568a>.